

- ANSELL, G. B. & KATZ, L. (1966). *Acta Cryst.* **21**, 482.
 CANNERI, G. (1925). *Gazz. Chim. Ital.* **55**, 611.
 COTTON, F. A. & WOOD, J. S. (1964). *Inorg. Chem.* **3**, 245.
 DOYLE, P. A. & TURNER, P. S. (1968). *Acta Cryst.* **A24**, 390.
 FURNAS, T. C. (1957). *Single Crystal Orienter Instruction Manual*. Milwaukee: General Electric Company.
 HAAS, D. J., HARRIS, D. R. & MILLS, H. H. (1965). *Acta Cryst.* **19**, 676.
 HOUSTY, J. & CLASTRE, J. (1957). *Acta Cryst.* **10**, 695.
 MONTGOMERY, H., CHASTAIN, R. V., NATT, J. J., WITKOWSKA, A. M. & LINGAFELTER, E. C. (1967). *Acta Cryst.* **22**, 775.
 STEWART, J. M. (1964). *Crystal Structure Calculations System X-ray-63*, Univ. of Maryland TR-64-6.
 STEWART, R. F., DAVIDSON, E. R. & SIMPSON, W. T. (1965). *J. Chem. Phys.* **42**, 3175.

Acta Cryst. (1970). **B26**, 341

Investigations of Dithienylglycollic Esters. II. The Crystal and Molecular Structure of Quinuclidinyl Di- α,α' -thienylglycolate

BY ANITA MEYERHÖFFER

Department of Medical Physics, Karolinska Institutet, S-104 01 Stockholm and The Research Institute of National Defence, S-172 04 Sundbyberg, Sweden

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Crystals of quinuclidinyl di- α,α' -thienylglycolate are monoclinic, space group $P2_1$ with two formula units in a cell having the dimensions $a = 9.762$, $b = 11.348$, $c = 8.556 \text{ \AA}$, $\beta = 118.405^\circ$. The crystal structure was solved by a three-dimensional Patterson synthesis and refined by three-dimensional least-squares to a final R value of 0.042. The two thiophene rings make an angle of 96° to each other. The glycolic bridge connecting the thiophene rings with the quinuclidine is almost planar. The quinuclidine skeleton is twisted 5° around its threefold axis. An intermolecular hydrogen bond is present between the hydroxyl hydrogen atom of the glycolic bridge and the nitrogen atom in the quinuclidine.

Introduction

The crystal structure of the anticholinergic drug, quinuclidinyl di- α,α' -thienylglycolate,



has been solved as part of our programme on the correlation between molecular configuration and pharmacological action of psychotomimetic compounds. The crystal structure of the closely related quinuclidinyl benzilate has been determined in the hydrobromide salt (Meyerhöffer & Carlström, 1969). As one of the important factors in psychotomimetic drugs is probably the availability of a free electron pair on a nitrogen atom (Gabel & Abood, 1965), it was regarded as important to solve the structure of a psychoactive base. All attempts to determine the structure of the base of the benzilate proved unsuccessful. For this reason the equally active thiophene compound was chosen as it could be analyzed directly by the 'heavy-atom' method. Another advantage was that the hydrogen positions could be determined, which was not possible in the case of the hydrobromide salt because of the strong interference of the heavy bromine atoms.

Experimental

The synthesis of quinuclidinyl di- α,α' -thienylglycolate has been described by Nyberg, Östman & Wallerberg

(1970). Large, colourless, optically perfect crystals were prepared through recrystallization in acetone during slow cooling. The unit-cell dimensions were determined from Weissenberg photographs and from diffractometer measurements. The density of single crystals was established by flotation in mixed solvents.

Crystal data

Quinuclidinyl di- α,α' -thienylglycolate, $\text{C}_{17}\text{H}_{19}\text{NO}_3\text{S}_2$
 F. W. 349.47

$$\begin{aligned} a &= 9.762 \pm 0.005 \text{ \AA} \\ b &= 11.348 \pm 0.005 \text{ \AA} \\ c &= 8.556 \pm 0.005 \text{ \AA} \\ \beta &= 118.405^\circ \pm 0.010^\circ \\ V &= 833.72 \text{ \AA}^3 \\ D_m &= 1.393 \pm 0.005 \text{ g.cm}^{-3} \\ D_x &= (Z=2) 1.392 \pm 0.002 \text{ g.cm}^{-3} \end{aligned}$$

The only systematic absences were $0k0$ when k is odd. Since the molecule is asymmetric and the unit cell contains two molecules the space group can only be $P2_1$.

Intensity data were collected with a Philips automatic single-crystal diffractometer (PAILRED) using monochromatic $\text{Mo K}\alpha$ radiation. The crystal, having the dimensions $0.42 \times 0.62 \times 0.41$ mm, was oriented about the b axis and 15 layers corresponding to the entire Cu sphere were recorded. There were 2171 independent reflexions of which 146 were too weak to be

observed. The unobserved reflexions were given half the intensity of the threshold value. The intensities were corrected for corresponding background, $\alpha_1-\alpha_2$ separation, Lorentz and polarization factors, but not for extinction or absorption. The structure amplitudes were placed approximately on an absolute scale by Wilson statistics. The above calculations were performed with the program system of Ringertz (1969). The X-ray 63 program system (Stewart & High, 1965) was used during the following procedures except for the perspective drawing, which was computed with the ORTEP program (Johnson, 1965). All calculations were carried out on an IBM 7090 computer.

Determination of the structure

A three-dimensional Patterson function was calculated from the corrected set of diffractometer data. In the Harker section $y=\frac{1}{2}$, two heavy peaks were located. All four combinations of these two peaks were tried as S-S rotation peaks, but calculations of structure amplitudes showed essentially no differences in the reliability indices R ($R = \sum |F_o - F_c| / \sum |F_c|$) between the different alternatives. Instead a combination of one of

the heavy peaks with another peak having roughly half the intensity of the first chosen caused a marked drop in the R value, indicating that the true S positions had been found. The electron density map, based upon the sulphur signs only, showed, together with the sulphur maxima, additional peaks in the vicinity of these maxima and one of the five-membered rings could be identified. Iterated Fourier refinement gave the other five-membered ring and finally the entire molecule. The R value at this stage was 0.19 with an overall temperature factor $B=3.5 \text{ \AA}^2$. It was now possible to interpret correctly the Patterson function; one of the two heavy peaks in the Harker section $y=\frac{1}{2}$ was found to be composed of three coincident vectors, namely O-O, S-O and S-C.

Refinement procedure

The initial refinement was carried out by a block-diagonal three-dimensional least-squares program because of the large number of parameters involved. However, for the final refinement a full-matrix program was used. In this case the refinement had to be carried out in two steps. Initially only the strongest 1650 of the 2171 reflexions were used and the structure amplitudes were

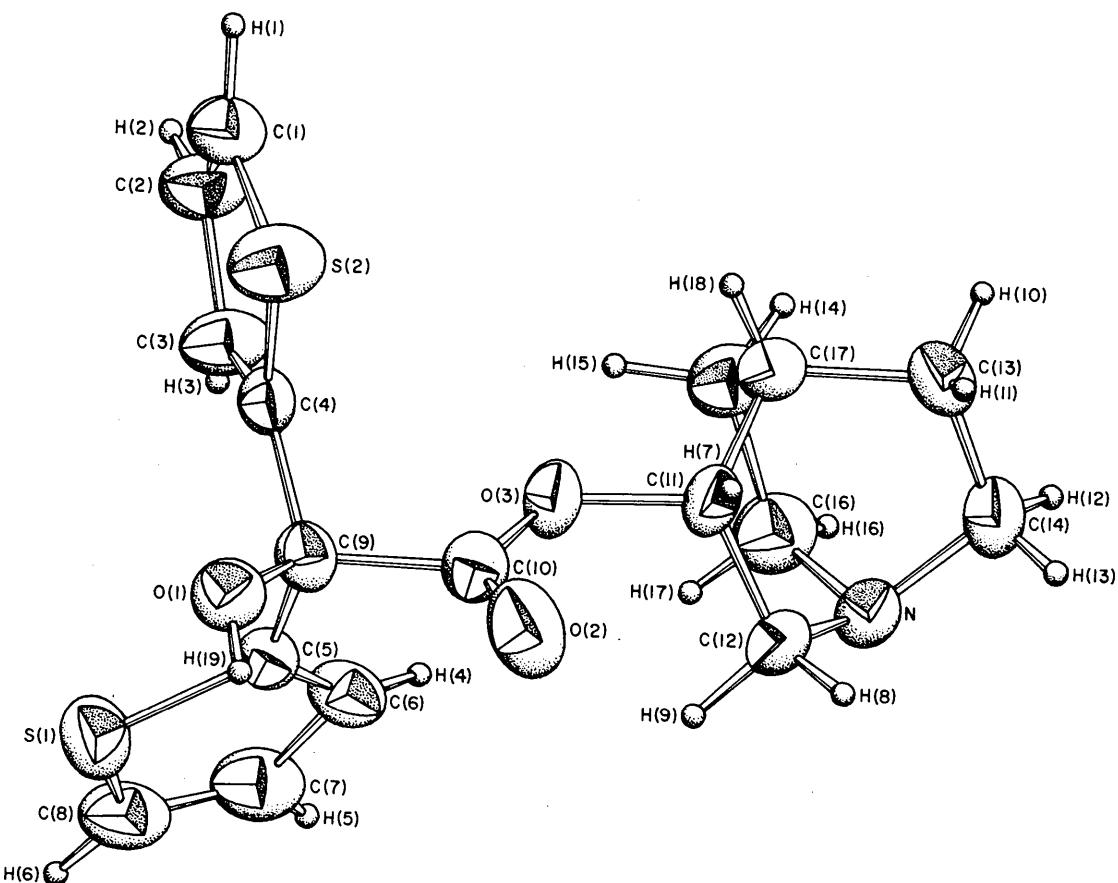


Fig. 1. A perspective drawing of the quinuclidinyl di- α,α' -thienylglycolate molecule, showing the ellipsoids of thermal motion with a probability of 50%.

Table 1. Observed and calculated structure factors ($\times 10$).

Asterisks indicate unobserved reflexions

H	K	L	F _G	F _C	A _C	B _C	H	K	L	F _G	F _C	A _C	B _C	H	K	L	F _G	F _C	A _C	B _C	H	K	L	F _G	F _C	A _C	B _C		
13	0	-5	20	12	-12	0	10	6	-1	70	77	-37	-67	9	7	-1	73	70	65	65	5	10	-L	82	86	-8L	0		
12	6	-5	29	28	-26	0	10	6	-2*	70	71	-6	-9	9	2	-2	103	103	51	9C	7	13	-3	26	23	23	2	6	
12	5	-8	15	27	-24	-13	10	6	-1	32	31	-31	-1	9	2	-1	62	63	38	56	7	12	-6	53	48	21	0	6	
12	5	-7	23	26	-20	17	10	5	-10	32	31	-28	-14	9	1	-1	94	24	3	-24	7	12	-5	35	33	-25	-21	6	
12	5	-6	49	44	-34	29	10	5	-9	12	30	79	9	9	1	-2	33	35	-29	-26	7	12	-3	23	20	17	-23	6	
12	5	-5	39	39	-35	35	10	5	-8	17	10	7	9	9	1	-2	51	52	-43	-40	7	12	-2	20	17	-19	-19	6	
12	5	-4	34	34	-15	-31	10	5	-7*	10	7	16	4	9	1	-2	24	22	-2	-5	7	12	-1	3	29	-27	-11	-11	6
12	5	-3	43	44	9	-43	10	5	-6	40	52	-3	-2	9	1	-2	92	93	-91	-18	7	11	-7	34	28	4	28	6	
12	5	-2	7	14	-4	-1	10	5	-5*	7	6	-4	5	5	1	-2	63	62	-62	-5	7	11	-6	45	39	34	34	6	
12	4	-5	30	5	-29	0	10	5	-4	20	30	-22	-2	9	1	-2	32	32	-22	-2	7	11	-5	37	37	37	37	6	
12	7	-2	76	73	73	0	10	5	-5	3	34	30	-22	-21	9	1	-4	73	66	-46	-6	7	11	-4	29	26	13	22	6
12	4	-6	33	34	19	-28	10	5	-5	2	55	52	-38	-49	9	1	-3	75	75	-60	-75	7	11	-3	29	26	13	-23	6
12	4	-5	67	65	29	-58	10	5	-5	1	46	46	-43	-17	9	1	-2	45	46	-9	-46	7	11	-2	33	32	-57	-59	6
12	4	-4	42	42	42	-42	10	5	-4	10	10	20	14	9	1	-2	51	51	-51	-51	7	11	-1	71	71	14	14	6	
12	3	-5	17	17	11	-13	10	4	-9	23	27	12	-24	9	0	-1	64	51	-53	0	7	10	-7	15	17	15	15	6	
12	4	-2	21	14	-14	1	10	4	-8	28	51	47	-20	9	C	-10	7	20	-20	0	7	10	-7	32	37	34	-12	5	
12	3	-7	59	61	55	-58	10	4	-6	59	65	58	-31	9	C	-9	29	20	-20	0	7	10	-6	42	39	-24	31	5	
12	3	-7	40	37	14	-31	10	3	-7	32	26	-26	-1	6	12	-5	54	54	-54	-54	7	9	-6	50	61	55	17	6	
12	3	-5	35	35	15	-35	10	4	-5	87	86	86	-34	9	1	-2	56	56	-56	-56	7	10	-5	45	45	-45	-45	6	
12	3	-5	15	14	-14	-9	10	3	-4	22	30	22	-2	9	0	-6	16	10	-10	l	7	10	-1	26	25	-23	-10	6	
12	3	-4	7	10	8	6	10	4	-3	24	23	20	11	9	C	-5	81	80	-80	0	7	10	-2	39	40	22	-23	6	
12	3	-3	17	9	6	-1	10	4	-2	20	19	10	-15	9	C	-4	132	132	-132	D	7	10	-1	40	31	-31	-21	6	
12	2	-5	26	26	17	19	10	3	-10	34	31	-8	-30	9	0	-2	17	18	-18	0	7	9	-8	32	32	-23	-22	6	
12	2	-8	66	61	7	66	10	3	-10	34	31	-8	-30	9	0	-2	121	121	-121	D	7	9	-9	60	61	-55	5	6	
12	2	-6	37	40	-17	38	10	3	-7	32	26	-26	-1	6	12	-5	7	15	-13	-7	7	9	-6	47	49	-26	-26	6	
12	2	-5	16	16	9	-16	10	3	-5	42	43	-43	-1	8	12	-3	16	19	-12	-15	7	9	-9	43	43	-43	-5	5	
12	2	-3	14	14	-9	10	3	-4	35	34	4	-34	8	12	-2	36	27	16	-22	7	9	-3	39	36	-35	-35	6		
12	2	-2	57	55	11	54	10	3	-3	40	41	40	-10	8	11	-7	38	36	0	35	7	9	-2	55	52	-61	-61	6	
12	1	-8	25	23	12	-20	10	3	-1	30	32	26	-2	9	10	-1	20	19	10	-20	7	9	-6	55	52	-61	-61	6	
12	1	-5	27	26	17	19	10	2	-10	49	51	51	-49	9	11	-4	32	27	2	-27	7	9	-8	64	64	-64	-64	6	
12	1	-4	20	29	9	-9	10	2	-9	63	63	55	-52	8	11	-3	40	41	16	-17	7	9	-7	61	61	-54	-54	6	
12	1	-2	80	66	7	66	10	2	-7	56	57	57	-54	6	11	-2	45	44	-30	-31	7	9	-6	52	50	-52	-52	6	
12	1	-2	14	14	-9	10	1	-2	25	27	26	-6	8	10	-7	59	65	65	-55	7	9	-8	43	41	-41	-41	6		
12	0	-8	17	18	-16	0	10	2	-5	67	61	52	-52	8	10	-6	20	16	-2	20	7	9	-8	80	79	6	6	6	
12	0	-7	67	67	-58	0	10	2	-4	50	49	36	-34	8	10	-6	20	16	-2	20	7	9	-8	66	64	-66	-66	6	
12	0	-5	34	34	-16	0	10	2	-4	25	24	23	-2	8	10	-6	19	19	-1	19	7	9	-8	55	53	-55	-55	6	
12	0	-4	34	34	-16	0	10	2	-2	23	27	27	-4	10	14	-2	22	19	15	-15	7	9	-8	52	52	-52	-52	6	
12	0	-4	14	14	-9	10	1	-2	25	28	26	-10	8	10	-1	20	19	15	-15	7	9	-8	52	52	-52	-52	6		
12	0	-3	26	27	-26	8	10	1	-1	35	34	18	-18	9	2	-2	31	31	0	0	7	9	-9	108	108	-108	-108	6	
11	7	-5	26	11	-11	-29	10	9	-1	35	34	18	-18	9	2	-2	31	31	-18	-18	7	9	-9	108	108	-108	-108	6	
11	7	-4	24	24	-12	1	10	9	-1	30	29	18	-8	9	2	-2	31	30	-18	-18	7	9	-9	108	108	-108	-108	6	
11	7	-2	42	42	25	20	10	9	-9	0	27	21	21	0	9	2	-2	50	51	-51	9	7	9	-9	154	154	-154	-154	6
11	6	-8	36	37	-25	-30	10	9	-8	85	89	89	-80	9	7	-8	39	37	-38	0	7	6	-6	64	66	-54	-54	6	
11	6	-7	36	37	-25	-30	7	9	-7	36	38	38	-30	8	6	-7	26	26	-26	0	7	6	-6	64	66	-54	-54	6	
11	6	-6	81	79	79	-79	7	9	-6	36	38	38	-30	8	6	-6	26	26	-26	0	7	6	-6	64	66	-54	-54	6	
11	6	-5	80	79	79	-79	7	9	-5	36	38	38	-30	8	6	-5	18	18	-18	0	7	6	-6	64	66	-54	-54	6	
11	6	-4	40	38	-26	-28	7	9	-4	38	40	40	-40	8	6	-4	28	28	-28	0	7	6	-6	64	66	-54	-54	6	
11	6	-3	39	39	-23	-21	7	9	-3	38	40	40	-40	8	6	-3	28	28	-28	0	7	6	-6	64	66	-54	-54	6	
11	5	-8	35	34	-24	-23	9	8	-6	55	57	57	-54	8	7	-5	33	33	-33	0	7	6	-6	64	66	-54	-54	6	
11	5	-7	34	34	-24	-23	9	8	-5	56	57	57	-54	8	7	-5	33	33	-33	0	7	6	-6	64	66	-54	-54	6	
11	5	-6	34	34	-24	-23	9	8	-4	56	57	57	-54	8	7	-4	33	33	-33	0	7	6	-6	64	66	-54	-54	6	
11	5	-5	34	34	-24	-23	9	8	-3	56	57	57	-54	8	7	-3	33	33	-33	0	7	6	-6	64	66	-54	-54	6	
11	5	-4	34	34	-24	-23	9	8	-2	56	57	57	-54	8	7	-2	33	33	-33	0	7	6	-6	64	66	-54	-54	6	
11	5	-3	34	34	-24	-23	9	8	-1	56	57	57	-54	8	7	-1	33	33	-33	0	7	6	-6	64	66	-54	-54	6	
11	5	-2	34	34	-24	-23	9	8	0	56	57	57	-54	8	7	0	33	33	-33	0	7	6	-6	64	66	-54	-54	6	
11	5	-1	34	34	-24	-23	9	8	-1	56	57	57	-54	8	7	-1	33	33	-33	0	7	6	-6	64	66	-54	-54	6	
11	4	-8	23	23	22	22	10	9	-6																				

INVESTIGATIONS OF DITHIENYLGLYCOLIC ESTERS. II

Table 1 (cont.)

H	K	L	FO	FC	AC	BC	H	K	L	FO	FC	AC	BC	H	K	L	FO	FC	AC	BC	H	K	L	FO	FC	AC	BC								
5	9	-5	57	61	13	40	4	8	-5	100	102	-27	-99	3	7	-7	63	64	-60	23	2	6	-3	126	119	-92	75	1	4	-4	159	159	-165	-35	
5	9	-4	70	75	67	34	4	8	-4	120	122	-40	-115	3	7	-6	162	163	-38	-54	2	6	-1	74	69	-111	-42	1	4	-2	367	366	-356	362	
5	9	-2	33	31	-4	31	4	8	-2	38	37	30	-22	3	7	-5	79	81	-63	-50	2	5	-10	17	16	-3	-15	1	4	-1	358	377	-369	78	
5	9	-1	102	105	13	104	4	8	-1	107	103	-25	-17	3	7	-4	132	133	-73	-96	2	5	-9	45	42	-37	-16	1	3	-10	211	207	-21	27	
5	9	-1	9	35	34	17	4	8	-1	102	107	-25	-17	3	7	-3	120	116	-74	-74	2	5	-8	50	53	-17	-16	1	3	-9	62	45	-43	3	
5	8	-8	7	8	-8	-1	4	7	-9	44	48	-33	-35	3	7	-1	66	62	-21	-21	2	5	-7	14	9	-6	-8	1	3	-8	58	58	-5	-58	
5	8	-7	60	59	57	-15	4	7	-8	34	36	36	-8	3	6	-10	92	92	5	42	2	5	-6	117	115	-108	-108	1	3	-7	69	90	86	26	
5	8	-6	16	19	-19	2	4	7	-7	65	66	-62	-21	3	6	-5	100	98	-12	-21	2	5	-5	60	46	-36	31	1	3	-5	128	140	-114	-14	
5	8	-5	60	64	-14	2	4	7	-5	32	30	-21	-18	3	6	-4	127	127	-27	-27	2	5	-4	149	149	-148	-148	1	3	-4	122	122	-122	-122	
5	8	-4	63	61	-31	-52	4	7	-4	116	143	-11	-144	3	6	-3	112	114	-102	-51	2	5	-3	199	192	-178	-176	1	3	-3	215	214	-159	-203	
5	8	-3	72	72	-25	-68	4	7	-4	116	143	-11	-144	3	6	-2	105	103	-30	-25	2	5	-2	105	103	-32	-97	1	3	-3	507	495	-84	407	
5	8	-2	103	103	52	90	4	7	-3	15	17	-26	-16	3	6	-1	131	132	-13	-13	2	5	-1	61	57	-32	-47	1	3	-2	53	101	-101	-95	
5	8	-1	103	103	10	110	4	7	-2	131	132	-16	-10	3	6	-1	166	166	-16	-16	2	5	-1	133	132	-111	-111	1	3	-1	117	116	-116	-112	
5	7	-10	35	31	31	-2	4	7	-1	22	23	-23	-9	3	6	-1	217	211	-11	-11	2	5	-9	22	18	-14	18	1	2	-10	22	26	-23	23	
5	7	-9	50	52	-41	-32	4	6	-10	7	12	-3	-12	3	6	-1	126	124	-13	-12	2	5	-8	95	94	-81	46	1	2	-9	56	56	-6	-56	
5	7	-8	32	29	-22	-19	4	6	-9	56	52	-51	-6	3	6	-1	126	124	-119	-159	2	5	-7	7	14	-22	13	1	2	-8	57	50	-80	-42	
5	7	-7	31	31	-2	-1	4	6	-8	51	52	-53	-32	3	6	-1	126	124	-119	-159	2	5	-6	15	15	-15	-15	1	2	-6	128	128	-128	-128	
5	7	-6	55	52	-41	-32	4	6	-7	71	73	-33	-33	3	6	-5	102	102	-22	-19	2	5	-4	212	211	-108	181	1	2	-5	43	48	-7	-47	
5	7	-5	64	67	-67	-2	4	6	-6	38	39	-33	-39	3	6	-4	112	114	-21	-51	2	5	-3	216	212	-25	211	1	2	-4	180	177	52	-169	
5	7	-4	210	211	-80	-195	4	6	-5	93	100	-79	-116	3	6	-3	111	112	-21	-91	2	5	-2	289	281	-32	31	1	2	-3	242	227	209	-91	
5	7	-3	203	201	-41	-197	4	6	-2	101	102	-37	-37	3	6	-1	156	153	-57	-57	2	5	-1	196	201	-161	120	1	2	-2	308	270	-170	-257	
5	7	-2	99	99	1	-99	4	5	-2	91	91	64	64	3	5	-1	164	163	-152	-57	2	5	-1	31	31	4	31	1	2	-1	207	670	-339	683	
5	6	-16	16	19	-8	17	4	6	-1	62	59	-53	9	3	5	-3	157	153	-117	-53	2	5	-2	10	10	-7	-8	1	1	-10	22	19	-5	19	
5	6	-15	29	28	-24	1	4	5	-1	101	117	-23	-18	3	5	-1	172	172	-63	-63	2	5	-1	52	52	-50	-50	1	2	-9	60	62	-52	33	
5	6	-14	19	19	-2	21	4	5	-1	10	32	-32	33	3	5	-1	10	32	-32	2	5	-1	92	95	-93	-93	1	2	-8	61	62	-61	33		
5	6	-13	69	66	19	19	4	5	-9	44	46	44	-44	3	5	-10	11	26	26	2	5	-8	11	11	-71	-87	1	1	-7	116	111	110	15		
5	6	-12	113	111	-62	-116	4	5	-8	53	55	34	45	3	5	-7	125	125	-62	-62	2	5	-6	155	155	-153	-153	1	2	-5	308	270	-270	-257	
5	6	-11	133	133	33	131	4	5	-7	194	187	-87	-173	3	5	-6	124	125	-15	-104	2	5	-5	10	10	-14	-23	1	1	-1	292	296	166	-246	
5	6	-10	35	38	-35	-15	4	5	-5	171	176	174	24	3	5	-4	128	129	-165	-165	2	5	-3	21	21	-14	-23	1	1	-1	214	212	212	0	
5	6	-9	50	49	1	-50	4	5	-4	101	102	51	48	3	5	-3	101	102	306	306	2	5	-2	207	206	-110	338	1	1	-1	165	151	-151	0	
5	6	-8	25	26	-23	17	4	5	-10	43	40	-40	2	3	5	-1	130	131	-224	-123	2	5	-1	137	135	-135	-135	1	2	-7	167	165	-168	168	
5	6	-7	47	47	44	18	4	5	-6	11	15	-10	-10	3	5	-1	104	105	-32	-27	2	5	-5	126	125	-118	-118	1	2	-6	154	154	-154	154	
5	6	-6	51	50	-20	20	4	5	-5	112	111	-103	-103	3	5	-4	106	107	-32	-27	2	5	-4	205	204	-199	-199	1	2	-5	215	220	-220	0	
5	6	-5	171	174	172	-26	4	5	-4	7	52	51	-50	-43	3	5	-3	101	103	-65	-65	2	5	-3	213	212	-152	-152	1	2	-2	338	338	0	0
5	6	-4	113	120	116	30	4	5	-3	190	193	-169	-26	3	5	-2	106	107	-30	-25	2	5	-2	305	305	-305	-305	1	2	-1	605	603	-603	603	
5	6	-3	120	120	116	41	4	5	-2	247	240	-240	-26	3	5	-1	130	140	-68	-123	2	5	-1	207	207	-120	-120	1	2	-1	520	516	-514	440	
5	6	-2	120	120	116	41	4	5	-1	229	224	-223	-123	3	5	-1	124	124	-127	-126	2	5	-1	207	207	-120	-120	1	2	-1	495	493	-493	493	
5	6	-1	110	110	30	30	4	5	-1	361	363	92	35	3	5	-1	142	142	-452	-452	2	5	-1	452	453	-453	-453	1	2	-1	133	132	-127	-127	
5	5	-9	74	77	60	-48	4	5	-11	16	8	-8	-1	3	5	-11	7	17	17	-17	2	5	-1	205	202	-202	-202	1	2	-9	229	229	-83	-83	
5	5	-8	94	94	39	-29	4	5	-10	11	19	19	-2	3	5	-10	11	31	31	-22	2	5	-1	205	204	-156	-156	1	2	-8	156	156	-156	-156	
5	5	-7	89	89	39	-29	4	5	-9	19	19	19	-2	3	5	-9	19	21	21	-22	2	5	-1	205	204	-156	-156	1	2	-7	205	204	-156	-156	
5	5	-6	89	89	39	-29	4	5	-8	19	19	19	-2	3	5	-8	19	21	21	-22	2	5	-1	205	204	-156	-156	1	2	-6	156	156	-156	-156	
5	5	-5	89	89	39	-29	4	5	-7	19	19	19	-2	3	5	-7	19	21	21	-22	2	5	-1	205	204	-156	-156	1	2	-5	205	204	-156	-156	
5	5	-4	89	89	39	-29	4	5	-6	19	19	19	-2	3	5	-6	19	21	21	-22	2	5	-1	205	204	-156	-156	1	2	-4	205	204	-156	-156	
5	5	-3	89	89	39	-29	4	5	-5	19	19	19	-2	3	5	-5	19	21	21	-22	2	5	-1	205	204	-156	-156	1	2	-3	205	204	-156	-156	
5	5	-2	89	89	39	-29	4	5	-4	19	19	19	-2	3	5	-4	19	21	21	-22	2	5	-1	205	204	-156	-156	1	2	-2	205	204	-156	-156	
5	5	-1	89	89	39	-29</																													

Table 1 (cont.)

H	K	L	F _u	F _C	A _C	B _C	H	K	L	F _u	F _C	A _C	B _C	H	K	L	F _u	F _C	A _C	B _C	H	K	L	F _u	F _C	A _C	B _C							
0	10	6	29	28	-17	22	1	13	2	21	20	19	8	3	1	/	56	59	-10	-58	4	6	4	45	45	-13	-15	6	1	2	77	77	7	76
0	10	1	28	45	45	-20	1	13	4	47	41	41	-8	3	2	0	34	40	-40	-4	4	6	5	34	29	-27	-11	6	1	4	25	15	-11	-47
0	11	2	47	49	8	-48	1	14	0	31	36	23	-27	3	2	1	143	125	-56	-138	4	6	7	41	42	-19	-38	6	1	5	49	47	-22	-42
0	11	3	47	46	-10	-45	1	14	1*	7	12	-12	-8	3	2	2	36	31	-169	-169	4	6	7	70	95	-97	-24	6	1	6	27	31	11	29
0	11	4	51	50	-10	-23	1	14	2	21	21	17	-10	3	2	3	12	12	-111	-111	4	7	1	105	107	-107	-155	6	2	0	153	133	-134	-75
0	11	5	51	32	-30	-12	1	14	3*	7	19	-18	-3	3	2	4	19	12	-30	-31	4	7	1	105	107	-107	-155	6	2	2	153	133	-134	-75
0	11	6	46	44	2	44	1	15	0	16	27	-5	-26	3	2	5	17	11	-10	-10	4	7	4	50	51	-42	-27	6	2	3	69	73	-65	33
0	11	7	31	35	25	-24	2	0	0	448	462	462	0	3	2	6	20	19	-30	-31	4	7	5	40	41	-37	-22	6	2	5	60	60	-53	30
0	12	0	20	26	-15	-25	2	1	1	20	16	16	0	3	2	4	19	12	-30	-31	4	7	5	40	44	-37	-22	6	2	6	60	60	-53	30
0	12	1	26	59	-13	-13	2	0	1	20	38	38	0	3	2	4	19	12	-30	-31	4	7	5	40	44	-37	-22	6	2	6	60	60	-53	30
0	12	2	24	22	17	14	2	0	3	104	102	-102	0	3	1	0	1	19	-156	114	114	8	0	89	87	-85	-17	6	2	6	19	21	-21	14
0	12	3	35	25	-22	11	2	0	4	15	12	18	0	3	1	1	111	99	99	49	4	8	1	96	92	-84	-36	6	3	0	153	155	-122	96
0	12	4	32	31	17	35	2	0	5	83	-83	0	0	3	1	1	21	25	-79	29	4	8	2	34	33	-31	-11	6	3	0	93	89	-87	-15
0	12	5	32	32	-2	2	0	6*	11	11	0	0	3	3	4	192	192	-2	196	4	8	4	12	9	-9	-3	6	3	3	48	48	-44	-36	
0	12	6	33	47	-7	46	0	0	7	83	87	87	0	3	3	4	58	-51	54	54	4	8	5*	7	7	3	3	48	46	-42	-36			
0	12	7	15	1	0	-17	2	0	8	26	31	0	0	3	3	4	51	-44	51	51	4	8	5*	14	14	-14	-16							
0	12	8	17	17	34	-34	2	0	9	25	25	25	0	3	3	4	25	-29	29	31	4	8	5*	14	14	-14	-16							
0	12	9	31	27	3	26	2	1	477	463	-463	-92	3	3	5	19	16	6	6	4	9	1	43	49	-47	-26								
0	12	10	17	15	6	-14	2	1	2	432	412	-375	-169	3	4	0	1	17	14	14	14	4	9	1	21	19	-19	-11						
0	12	11	17	15	6	-14	2	1	2	80	84	84	0	3	4	1	17	14	14	14	4	9	1	21	19	-19	-11							
0	12	12	45	47	45	-16	2	1	5	205	206	-63	-197	3	4	1	12	12	12	12	4	9	5	17	17	-28	-24							
0	12	13	7	15	13	-7	2	1	5	56	54	-49	-22	3	4	4	6	100	9	100	100	4	9	6	21	19	-17	-17						
0	12	14	7	12	12	-40	2	1	7	39	40	40	-54	3	4	4	6	15	-35	19	19	4	10	0	65	63	-57	-27						
0	12	15	20	21	21	-21	2	0	8	30	34	34	0	3	4	6	10	10	10	10	4	10	0	59	57	-52	-20							
0	12	16	455	506	506	0	2	1	9	13	17	12	-11	3	4	6	10	10	10	10	4	10	0	61	61	-52	-26							
0	12	17	7	63	83	83	0	2	2	0	331	331	141	-300	3	4	6	5	50	55	31	69	4	10	0	7	11	-2	-11					
0	12	18	7	63	83	83	0	2	2	105	105	105	0	3	4	6	5	50	55	31	69	4	10	0	60	60	-55	-25						
0	12	19	156	160	160	0	2	2	232	232	-218	-208	3	5	3	5	145	142	-83	-116	4	11	1	32	28	28	10							
0	12	20	31	31	31	0	2	2	4	63	58	-17	-55	3	5	3	5	145	142	-83	-116	4	11	1	32	28	28	10						
0	12	21	60	60	60	-29	2	2	5	68	76	-29	-70	3	5	4	63	63	-43	-66	4	11	2	57	59	-51	-27							
0	12	22	10	18	18	0	2	2	6	91	96	-28	-70	3	5	4	63	63	-43	-66	4	11	2	57	59	-51	-27							
0	12	23	19	19	19	0	2	2	6	105	105	105	0	3	5	4	63	63	-43	-66	4	11	2	57	59	-51	-27							
0	12	24	12	12	12	-29	2	2	3	145	150	12	147	3	5	6	63	63	-43	-66	4	11	2	57	59	-51	-27							
0	12	25	14	14	14	-24	2	2	3	60	65	65	0	3	5	6	63	63	-43	-66	4	11	2	57	59	-51	-27							
0	12	26	8	6	6	-6	2	2	40	45	-14	-42	3	6	7	35	38	-23	-1	4	14	0	42	42	-26	-32								
0	12	27	23	22	22	-2	2	2	3	7	23	23	-21	-11	3	7	0	125	129	-25	-127	5	0	0	159	160	-160	-160						
0	12	28	37	34	20	-27	2	2	7	52	52	-52	-31	3	7	0	136	132	-124	-124	5	0	0	173	177	-177	-177							
0	12	29	60	60	60	-29	2	2	7	52	52	-52	-31	3	7	0	136	132	-124	-124	5	0	0	173	177	-177	-177							
0	12	30	60	54	-54	-54	2	0	267	255	239	59	3	8	5	17	17	17	17	4	12	0	7	14	-14	-19								
0	12	31	450	436	259	-343	2	0	273	256	256	-28	3	7	4	27	37	-17	-85	5	0	0	88	86	-86	-86								
0	12	32	211	212	143	-157	2	0	281	261	199	-249	3	7	4	86	86	-86	-86	5	0	0	43	43	-38	-17								
0	12	33	31	30	30	-30	2	0	285	280	281	-166	3	8	5	27	27	-21	-21	5	0	0	130	132	-132	-132								
0	12	34	55	55	55	-37	2	0	286	286	286	0	3	8	5	27	27	-21	-21	5	0	0	100	99	-99	-99								
0	12	35	55	57	57	-37	2	0	287	287	287	0	3	8	5	27	27	-21	-21	5	0	0	130	132	-132	-132								
0	12	36	80	80	80	-31	2	0	75	75	75	-26	3	7	5	17	17	17	17	5	0	0	157	157	-157	-157								
0	12	37	80	80	80	-31	2	0	30	31	31	-15	3	7	5	17	17	17	17	5	0	0	167	165	-165	-165								
0	12	38	25	23	17	-16	2	0	84	83	83	-28	3	7	5	25	21	-10	-2	5	1	2	35	33	-33	-3								
0	12	39	132	127	127	-27	2	0	136	133	132	-12	3	7	5	25	20	-16	-1	5	1	2	21	20	-20	-20								
0	12	40	39	39	39	-39	2	0	181	181	181	-101	3	7	5	25	20	-16	-1	5	1	2	21	20	-20	-20								
0	12	41	103	103	103	-37	2	0	68	67	67	-6	3	7	5	25	20	-16	-1	5	1	2	21	20	-20	-20								
0	12	42	97	97	97	-37	2	0	58	57	57	-5	3	7	5	25	20	-16	-1	5	1	2	21	20	-20	-20								
0	12	43	97	97	97	-37	2	0	62	62	62	-1	3	7	5	25	20	-16	-1	5	1	2	21	20	-20	-20								
0	12	44	80	80	80	-31	2	0	75	75	75	-26	3	7	5	25	20	-16	-1	5	1	2	21	20	-20	-20								
0	12	45	123	123	123	-31	2	0	145	145	145	-31	3	7	5	25	20	-16	-1	5	1	2	21	20	-20	-20								
0	12	46	123	123	123	-31	2	0	151	152</td																								

given unit weight. The atomic scattering factors were those listed in *International Tables for X-ray Crystallography* (1962).

The first two cycles of block-diagonal refinement with non-hydrogen atoms only and isotropic temperature factors yielded an *R* value of 0.10. When anisotropic temperature factors were introduced the *R* value decreased to 0.075. A Fourier difference synthesis prepared at this stage made it possible to locate all of the 19 hydrogen atoms. With the hydrogen atoms included and given isotropic temperature factors, three cycles of refinement of the non-hydrogen atoms lowered the *R* value to 0.036. When also the positional parameters and the individual isotropic temperature factors of the hydrogen atoms were refined for one cycle the *R* value decreased to 0.033. The standard deviations of the atomic parameters were now quite satisfactory. In the next step the structure amplitudes were weighted. The weighting scheme adopted was $w = (a + bF_o + cF_o^2)^{-1}$ where *a*, *b* and *c* were derived from a quadratic curve of best fit for $|F_o - F_c|$ as a function of F_o . The numerical values of *a*, *b* and *c* were *a*=0.12, *b*=0.017, *c*=-0.0004 for $F_o \geq 10$ and *a*=0.282, *b*=*c*=0 for $F_o < 10$. A final refinement including all of the 2171 weighted structure factors resulted in *R*=0.042, i.e. somewhat higher than the lowest value for the 1650 reflexions, but the standard deviations of the atomic parameters were further decreased. The average of the shift to error ratio of the parameters was 0.20 for the non-hydrogen atoms and 0.85 for the hydrogen atoms. Observed and calculated structure factors from this last cycle of refinement are listed in Table 1. The positional and thermal parameters for the final structure are given in Tables 2 and 3 respectively.

Table 2. Final fractional atomic coordinates and their standard deviations (in parentheses)

	<i>x/a</i>	<i>y/b</i>	<i>z/c</i>
S(1)	0.0594 (1)	0.3500	1.0177 (1)
S(2)	0.0114 (1)	0.4893 (1)	0.4370 (1)
O(1)	0.1598 (3)	0.5242 (2)	0.8389 (3)
O(2)	0.4397 (4)	0.4686 (4)	0.8619 (5)
O(3)	0.3416 (3)	0.3041 (3)	0.7041 (3)
N	0.6342 (3)	0.0976 (3)	0.8229 (4)
C(1)	-0.1344 (5)	0.4053 (4)	0.2803 (5)
C(2)	-0.1562 (5)	0.3076 (4)	0.3496 (5)
C(3)	-0.0536 (4)	0.2979 (4)	0.5352 (4)
C(4)	0.0466 (4)	0.3899 (3)	0.6024 (4)
C(5)	0.1678 (4)	0.3181 (3)	0.9136 (4)
C(6)	0.2262 (5)	0.2028 (4)	0.9513 (5)
C(7)	0.1756 (6)	0.1463 (5)	1.0649 (5)
C(8)	0.0876 (5)	0.2132 (5)	1.1094 (5)
C(9)	0.1767 (4)	0.4104 (3)	0.7889 (4)
C(10)	0.3366 (4)	0.3995 (4)	0.7912 (4)
C(11)	0.4863 (4)	0.2835 (4)	0.6974 (5)
C(12)	0.6058 (4)	0.2171 (4)	0.8642 (5)
C(13)	0.6004 (5)	0.1804 (5)	0.5370 (5)
C(14)	0.7027 (5)	0.1058 (5)	0.7013 (6)
C(15)	0.3748 (5)	0.0912 (4)	0.5515 (5)
C(16)	0.4851 (5)	0.0345 (4)	0.7332 (6)
C(17)	0.4461 (4)	0.2060 (4)	0.5357 (5)
H(1)	-0.176 (6)	0.430 (5)	0.157 (7)
H(2)	-0.238 (7)	0.244 (6)	0.284 (8)

Table 2 (cont.)

H(3)	-0.058 (5)	0.239 (5)	0.582 (7)
H(4)	0.275 (4)	0.178 (4)	0.923 (5)
H(5)	0.215 (7)	0.069 (6)	1.099 (8)
H(6)	0.046 (6)	0.183 (5)	1.183 (7)
H(7)	0.522 (5)	0.360 (4)	0.684 (5)
H(8)	0.709 (4)	0.266 (4)	0.922 (5)
H(9)	0.565 (5)	0.210 (5)	0.956 (6)
H(10)	0.582 (5)	0.147 (5)	0.434 (6)
H(11)	0.651 (6)	0.256 (5)	0.549 (7)
H(12)	0.718 (6)	0.017 (5)	0.674 (7)
H(13)	0.802 (6)	0.138 (5)	0.765 (7)
H(14)	0.367 (7)	0.030 (5)	0.459 (8)
H(15)	0.268 (5)	0.117 (4)	0.541 (6)
H(16)	0.506 (7)	-0.043 (6)	0.722 (8)
H(17)	0.440 (5)	0.037 (4)	0.809 (6)
H(18)	0.379 (5)	0.254 (4)	0.438 (5)
H(19)	0.217 (6)	0.538 (5)	0.932 (6)

Description and discussion of the structure

The geometry of the molecule

The atomic nomenclature and the configuration of the molecule is illustrated in Fig. 1 which also shows the ellipsoids of thermal motion scaled to include 50 per cent probability. As can be seen, the molecule is composed of two planar α -substituted thiophene rings connected to a quinuclidine nucleus by means of an almost planar bridge, called the glycolic bridge in the following. The best planes through each of the two five-membered rings were calculated with and without the sulphur atoms included. In the former case all the atoms deviated significantly from the planes whereas only the sulphur atoms deviated significantly in the latter case. These deviations amounted to 0.010 Å for S(1) and 0.014 Å for S(2). This is slightly less than found in thenoic acid. Thus, Nardelli, Fava & Giraldi (1962) observed that the sulphur atom was displaced 0.028 Å from a plane through the carbon atoms in the ring of α -thenoic acid. Later, Hudson & Robertson (1962, 1964) confirmed this result. They found the

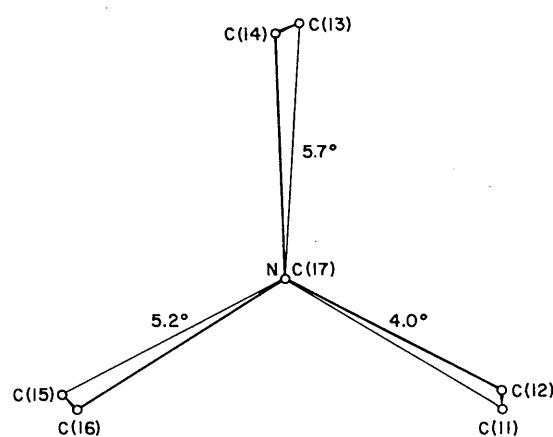


Fig. 2. The quinuclidine part of the molecule seen along its threefold axis, showing the angle of twist between the upper and lower halves.

deviation 0.02 Å of the sulphur atoms in α - and β -thenoic acids. The equations of the planes and the deviations of the atoms in the present structure are given in Table 4. The planes of the rings make an angle of 96.4° to each other.

The glycollic bridge connecting the thiophene rings to the quinuclidine part of the molecule is perfectly planar except for the hydroxyl oxygen atom, O(1), which deviates 0.34 Å from a plane defined by the atoms O(2), O(3), C(9), C(10), C(11). The equation of this plane and the atomic deviations are given in Table 4. The angles between the plane of the glycollic bridge and the thiophene rings C(1) to C(4) and C(5) to C(8) are 93° and 119° respectively.

The quinuclidine cage is almost symmetric in shape. However, when seen along its axis of symmetry,

C(17)...N, it exhibits a slight distortion illustrated in Fig. 2. The mean angle of twist is $5.0 \pm 0.7^\circ$ which is the same as that found in our earlier investigation of the quinuclidinyl benzilate (Meyerhöffer & Carlström, 1969). The question regarding the symmetry of quinuclidine and similar cage-like structures has already been discussed in this work. Recently, a new structure determination of a bicyclo[2.2.2]octane derivative has been reported by Cameron, Ferguson & Morris (1968) who likewise observed an angle of twist (3°) about the threefold axis of the carbocycle. According to these authors the distortion is adopted solely to release strain from torsional and nonbonded intramolecular interactions.

The overall geometry of the present molecule is very similar to that of the quinuclidinyl benzilate in spite

Table 3. Anisotropic parameters and standard deviations (in parentheses)

β_{ij} are the coefficients in the expression $\exp [-(\beta_{11}h^2 + \beta_{22}k^2 + \beta_{33}l^2 + 2\beta_{12}hk + 2\beta_{13}hl + 2\beta_{23}kl)]$.

	β_{11}	β_{22}	β_{33}	β_{12}	β_{13}	β_{23}
S(1)	0.0140 (1)	0.0099 (1)	0.0208 (2)	0.0004 (1)	0.0111 (1)	0.0007 (1)
S(2)	0.0186 (2)	0.0080 (1)	0.0183 (2)	-0.0019 (1)	0.0070 (2)	0.0030 (1)
O(1)	0.0130 (4)	0.0057 (2)	0.0159 (4)	0.0002 (2)	0.0057 (3)	-0.0023 (2)
O(2)	0.0153 (5)	0.0124 (4)	0.0289 (7)	-0.0066 (3)	0.0134 (5)	-0.0091 (5)
O(3)	0.0086 (3)	0.0064 (2)	0.0191 (5)	0.0004 (2)	0.0066 (3)	-0.0009 (3)
N	0.0098 (4)	0.0082 (3)	0.0144 (5)	0.0016 (3)	0.0051 (4)	0.0010 (3)
C(1)	0.0150 (6)	0.0100 (3)	0.0133 (6)	0.0013 (4)	0.0061 (5)	0.0007 (4)
C(2)	0.0146 (6)	0.0101 (4)	0.0135 (6)	-0.0025 (4)	0.0046 (5)	-0.0016 (4)
C(3)	0.0142 (5)	0.0068 (3)	0.0130 (5)	-0.0026 (3)	0.0037 (4)	0.0002 (3)
C(4)	0.0094 (4)	0.0055 (2)	0.0137 (5)	0.0008 (2)	0.0066 (4)	0.0002 (3)
C(5)	0.0093 (4)	0.0066 (3)	0.0112 (5)	-0.0005 (3)	0.0051 (4)	-0.0007 (3)
C(6)	0.0157 (6)	0.0067 (3)	0.0133 (6)	0.0006 (3)	0.0081 (5)	0.0014 (3)
C(7)	0.0185 (7)	0.0084 (4)	0.0156 (7)	-0.0009 (4)	0.0072 (6)	0.0019 (4)
C(8)	0.0152 (6)	0.0114 (5)	0.0132 (6)	-0.0039 (5)	0.0061 (5)	0.0010 (4)
C(9)	0.0102 (4)	0.0049 (2)	0.0143 (5)	0.0003 (3)	0.0065 (4)	-0.0007 (3)
C(10)	0.0097 (4)	0.0070 (3)	0.0144 (5)	-0.0010 (3)	0.0058 (4)	-0.0004 (3)
C(11)	0.0072 (4)	0.0075 (3)	0.0157 (6)	0.0005 (3)	0.0049 (4)	0.0005 (3)
C(12)	0.0085 (4)	0.0090 (3)	0.0132 (5)	0.0015 (3)	0.0032 (4)	-0.0001 (4)
C(13)	0.0144 (6)	0.0126 (5)	0.0167 (7)	0.0025 (4)	0.0102 (6)	0.0014 (5)
C(14)	0.0123 (5)	0.0120 (5)	0.0193 (7)	0.0036 (4)	0.0089 (5)	0.0011 (5)
C(15)	0.0119 (5)	0.0093 (4)	0.0162 (7)	-0.0010 (4)	0.0044 (5)	-0.0026 (4)
C(16)	0.0142 (6)	0.0086 (4)	0.0198 (8)	-0.0014 (4)	0.0072 (6)	0.0002 (5)
C(17)	0.0106 (5)	0.0099 (4)	0.0122 (5)	0.0019 (3)	0.0045 (4)	0.0010 (4)

Table 4. Equations of the best planes

The equations are of the form $m_1X + m_2Y + m_3Z = d$, X, Y, Z being the orthogonal coordinates in Å.

	Plane	m_1	m_2	m_3	d
A	C(1), C(2), C(3), C(4), S(2)	0.8587	-0.4743	-0.1941	-4.6989
B	C(1), C(2), C(3), C(4)	0.8563	-0.4787	-0.1939	-4.7087
C	C(5), C(6), C(7), C(8), S(1)	0.4995	0.3280	0.8018	5.6614
D	C(5), C(6), C(7), C(8)	0.4975	0.3308	0.8019	5.6733
E	C(9), C(10), C(11), O(2), O(3)	-0.0523	0.5289	0.8471	2.6423

Deviations (Å) from the best planes

Plane A	Plane B	Plane C	Plane D	Plane E
C(1) -0.002	C(1) 0.002	C(5) 0.004	C(5) 0.001	C(9) -0.002
C(2) -0.002	C(2) -0.004	C(6) -0.003	C(6) -0.001	C(10) -0.000
C(3) 0.006	C(3) 0.004	C(7) 0.000	C(7) 0.001	C(11) -0.002
C(4) -0.006	C(4) -0.002	C(8) 0.002	C(8) -0.001	O(2) 0.001
S(2) 0.004		S(1) -0.003		O(3) 0.003
	S(2) 0.014		S(1) -0.010	
				O(1) 0.344
				H(19) -0.152

of the latter being a salt. The main differences are due to twists around the C(4)–C(9) and C(5)–C(9) bonds connecting the rings to the glycollic bridge, a 23° twist around the C(9)–C(10) bond and a 20° twist around the O(3)–C(11) bond. A closer comparison of the two structures seems to be of little value because of the much lower accuracy in the determination of the benzilate structure.

Bond lengths and bond angles

The bond lengths and bond angles with estimated standard deviations are listed in Tables 5 and 6 and illustrated in Fig. 3 and 4. No correction has been made for thermal motion. The thiophene rings were non-symmetric when tested with the criteria of significance proposed by Cruickshank & Robertson (1953). Thus the C(1)–C(2) bond is significantly shorter than the C(3)–C(4) bond and the same is true for corresponding bonds in the other thiophene ring. C(1)–C(2) (1.321 Å) and C(7)–C(8) (1.331 Å) are moreover slightly shorter than the standard double bond length, of 1.337 Å. Such non-symmetric thiophene rings with short C–C bonds have earlier been described by Block, Filippakis & Schmidt (1967) in α -thienylacrylic acid.

Table 5. Interatomic distances and standard deviations (in parentheses)

S(2)–C(4)	1.712 (4) Å	C(11)–C(12)	1.543 (5) Å
S(2)–C(1)	1.708 (4)	C(13)–C(14)	1.535 (6)
C(1)–C(2)	1.321 (7)	C(15)–C(16)	1.548 (6)
C(2)–C(3)	1.422 (5)	C(11)–C(17)	1.524 (6)
C(3)–C(4)	1.358 (5)	C(13)–C(17)	1.529 (7)
		C(15)–C(17)	1.514 (7)
S(1)–C(5)	1.714 (5)	C(12)–N	1.461 (6)
S(1)–C(8)	1.702 (5)	C(14)–N	1.483 (7)
C(8)–C(7)	1.331 (8)	C(16)–N	1.469 (6)
C(7)–C(6)	1.434 (8)		
C(6)–C(5)	1.402 (6)	C(11)–H(7)	0.96 (5)
		C(12)–H(8)	1.04 (4)
C(1)–H(1)	0.98 (6)	C(12)–H(9)	1.04 (6)
C(2)–H(2)	1.02 (6)	C(13)–H(10)	0.89 (6)
C(3)–H(3)	0.79 (6)	C(13)–H(11)	0.96 (6)
C(6)–H(4)	0.69 (5)	C(14)–H(12)	1.06 (6)
C(7)–H(5)	0.94 (6)	C(14)–H(13)	0.93 (5)
C(8)–H(6)	0.96 (7)	C(15)–H(14)	1.03 (7)
		C(15)–H(15)	1.04 (5)
C(4)–C(9)	1.511 (4)	C(16)–H(16)	0.91 (7)
C(5)–C(9)	1.527 (6)	C(16)–H(17)	0.94 (6)
C(9)–C(10)	1.556 (6)	C(17)–H(18)	0.95 (4)
C(9)–O(1)	1.395 (5)		
C(10)–O(2)	1.190 (5)	O(1)···O(2)	2.719 (5)
C(10)–O(3)	1.329 (5)	O(1)···N'	2.752 (4)
O(3)–C(11)	1.460 (5)	H(19)···O(2)	2.63 (6)
O(1)–H(19)	0.74 (4)	H(19)···N'	2.02 (4)

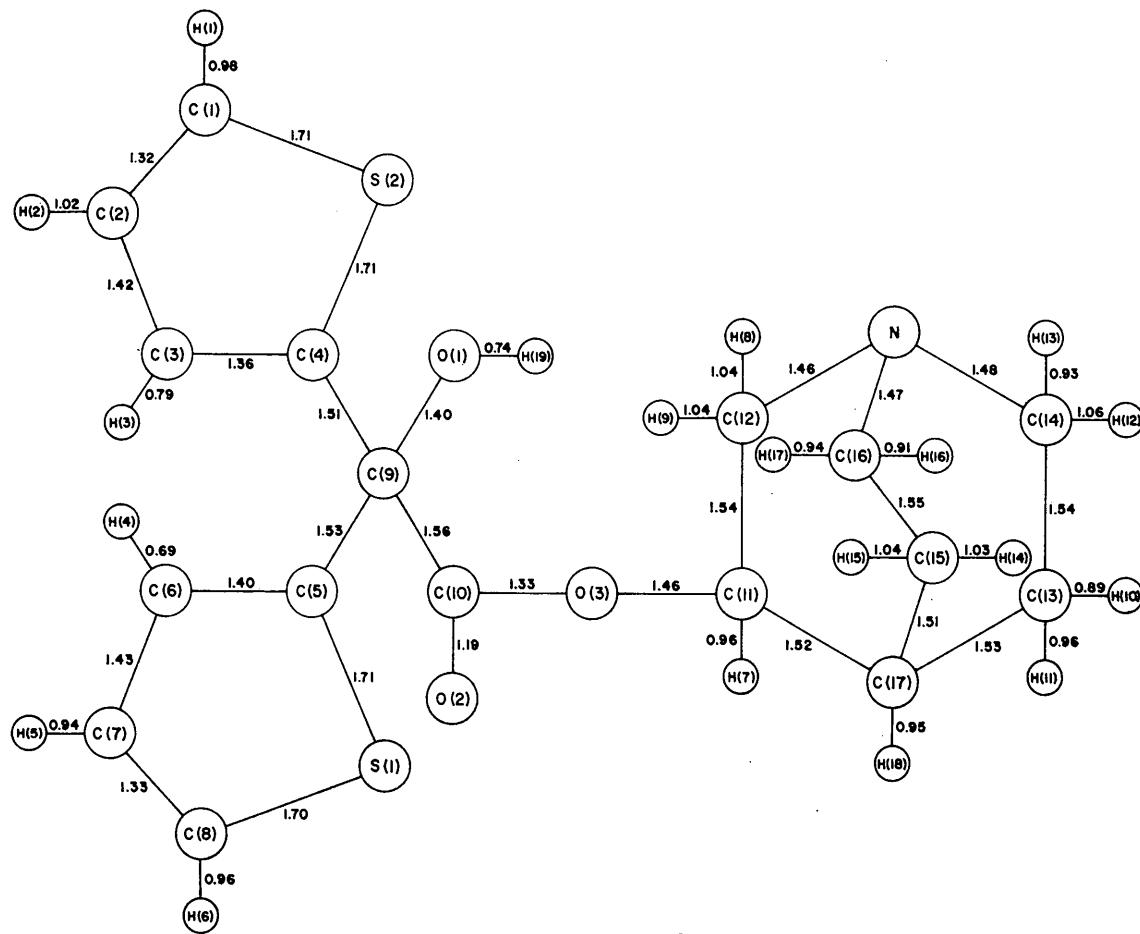


Fig. 3. Bond lengths (Å).

On comparing corresponding bond lengths in the two thiophene rings all bonds were found to be identical except for the C(3)–C(4) and C(5)–C(6) which were significantly different. Averaging the bond lengths of the thiophene rings, the mean S–C distance, 1.709 Å, is almost midway between a pure single bond, 1.815 Å and a pure double bond, 1.556 Å, as given by Bak, Christensen, Hansen-Nygaard & Rastrup-Andersen, (1961). The mean C–C bond length, 1.428 Å, opposite the sulphur atom and the mean C–C bond length, 1.353 Å, adjacent to the C–S bond deviate symmetrically with respect to the C–C bond length, 1.395 Å, of benzene. The mean bond lengths and the bond angles are all in excellent agreement with the earlier determinations of a α -substituted thiophene (Nardelli, Fava & Giraldi, 1962) as well as calculations of the thiophene structure from micro wave spectra (Bak, Christensen, Hansen-Nygaard & Rastrup-Andersen, 1961) and from electron diffraction (Momany, 1964). Large varia-

tions in bond lengths are reported for a β -substituted thiophene ring (Hudson & Robertson, 1964). On the other hand Visser, Heeres, Wolters & Vos (1968) have found normal thiophene bond lengths in α,α' - and β,β' -thienyls within the limits of the rather large experimental errors. In the rest of the molecule all bond lengths and bond angles are close to their normal values. The mean paraffinic C–C distance, 1.532 Å, the mean C–N distance, 1.471 Å and the mean C–H distance, 0.95 Å are all quite acceptable and so are the C=O and C–O bond lengths.

Hydrogen bonds and packing of the molecules

The observed planarity of the connecting glycollic bridge together with the short O(1)...O(2) distance, 2.719 Å, at first indicated the presence of an intramolecular hydrogen bond as presumed for the quinuclidinyl benzilate molecule (Meyerhöffer & Carlström, 1969). However, the short O(1)...N'(1-x, y+ $\frac{1}{2}$, 1-z)

Table 6. Interatomic angles and standard deviations (in parentheses)

C(4)–S(2)–C(1)	92.6 (2) $^\circ$	C(5)–S(1)–C(8)	92.4 (2) $^\circ$
S(2)–C(1)–C(2)	111.6 (3)	S(1)–C(8)–C(7)	111.9 (4)
C(1)–C(2)–C(3)	112.8 (4)	C(8)–C(7)–C(6)	114.6 (5)
C(2)–C(3)–C(4)	113.3 (4)	C(7)–C(6)–C(5)	109.9 (4)
C(3)–C(4)–S(2)	109.6 (2)	C(6)–C(5)–S(1)	111.2 (3)
C(3)–C(4)–C(9)	129.8 (3)	C(6)–C(5)–C(9)	130.4 (4)
S(2)–C(4)–C(9)	120.6 (3)	S(1)–C(5)–C(9)	118.2 (3)
S(2)–C(1)–H(1)	117 (3)	S(1)–C(8)–H(6)	127 (3)
C(2)–C(1)–H(1)	131 (3)	C(7)–C(8)–H(6)	121 (3)
C(1)–C(2)–H(2)	127 (4)	C(8)–C(7)–H(5)	132 (5)
C(3)–C(2)–H(2)	121 (4)	C(6)–C(7)–H(5)	114 (5)
C(2)–C(3)–H(3)	117 (3)	C(7)–C(6)–H(4)	125 (4)
C(4)–C(3)–H(3)	130 (3)	C(5)–C(6)–H(4)	125 (4)
C(4)–C(9)–C(5)	109.5 (3)	C(11)–C(12)–N	112.1 (3)
C(4)–C(9)–C(10)	109.6 (3)	C(13)–C(14)–N	111.3 (4)
C(5)–C(9)–C(10)	109.1 (3)	C(15)–C(16)–N	111.5 (4)
C(4)–C(9)–O(1)	107.7 (3)	C(11)–C(17)–C(13)	106.0 (3)
C(5)–C(9)–O(1)	111.5 (3)	C(13)–C(17)–C(15)	109.3 (4)
C(10)–C(9)–O(1)	109.5 (3)	C(15)–C(17)–C(11)	110.3 (4)
C(9)–C(10)–O(2)	123.7 (4)	C(12)–N–C(14)	108.2 (4)
C(9)–C(10)–O(3)	111.2 (3)	C(14)–N–C(16)	109.2 (3)
O(2)–C(10)–O(3)	125.1 (4)	C(16)–N–C(12)	108.9 (3)
C(10)–O(3)–C(11)	116.6 (3)	 	
O(3)–C(11)–C(12)	111.2 (4)	C(9)–O(1)–H(19)	113 (4)
O(3)–C(11)–C(17)	107.0 (3)	O(3)–C(11)–H(7)	106 (3)
C(12)–C(11)–C(17)	108.3 (3)	C(12)–C(11)–H(7)	113 (2)
C(14)–C(13)–C(17)	108.7 (4)	C(17)–C(11)–H(7)	112 (3)
C(16)–C(15)–C(17)	108.5 (3)	C(11)–C(12)–H(8)	108 (2)
N—C(12)–H(8)	111 (2)	C(17)–C(15)–H(14)	113 (4)
C(11)–C(12)–H(9)	110 (3)	C(16)–C(15)–H(15)	113 (3)
N—C(12)–H(9)	107 (3)	C(17)–C(15)–H(15)	104 (3)
H(8)–C(12)–H(9)	108 (4)	H(14)–C(15)–H(15)	114 (4)
C(14)–C(13)–H(10)	114 (3)	C(15)–C(16)–H(16)	112 (4)
C(17)–C(13)–H(10)	110 (3)	N—C(16)–H(16)	108 (4)
C(14)–C(13)–H(11)	108 (3)	C(15)–C(16)–H(17)	110 (2)
C(17)–C(13)–H(11)	106 (4)	N—C(16)–H(17)	107 (2)
H(10)–C(13)–H(11)	110 (6)	H(16)–C(16)–H(17)	108 (6)
C(13)–C(14)–H(12)	115 (3)	C(11)–C(17)–H(18)	104 (3)
N—C(14)–H(12)	105 (4)	C(13)–C(17)–H(18)	112 (3)
C(13)–C(14)–H(13)	111 (4)	C(15)–C(17)–H(18)	115 (3)
N—C(14)–H(13)	107 (4)	 	
H(12)–C(14)–H(13)	107 (5)	O(1)–H(19)…O(2)	89 (5)
C(16)–C(15)–H(14)	105 (3)	O(1)–H(19)…N'	173 (5)

distance, 2.752 Å, seemed equally well suited to an intermolecular hydrogen bond. When the position of H(19) was established it became evident that the H(19)...O(2) distance, 2.63 Å, was much too long to permit such a bond whereas the H(19)...N' distance, 2.02 Å was quite favourable. The existence of a hydrogen bond between O(1)H(19)...N' is further supported by the fact that the hydrogen atom lies almost on the straight line between the oxygen atom and the nitrogen atom, the O(1)-H(19)...N' angle being 173°. The hydrogen bond system is illustrated in Fig. 5.

The oxygen atom O(1) of the hydroxyl group is facing the sulphur atoms of the thiophene rings, the distances being as short as 2.939 Å and 3.056 Å. This verifies earlier findings that the van der Waals radius of sulphur is closer to 1.60 Å (Donohue, 1950; Nardelli, Fava & Giraldi, 1962), than to the value of 1.85 Å given by Pauling (1960) and even to the value of 1.72 Å supposed by Walker, Folting & Merritt (1969). All other distances are in accordance with ordinary requirements.

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References

- BAK, B., CHRISTENSEN, D., HANSEN-NYGAARD, L. & RASTRUP-ANDERSEN, J. (1961). *J. Mol. Spectr.* 7, 58.
 BLOCK, S., FILIPPakis, S. E. & SCHMIDT, G. M. J. (1967). *J. Chem. Soc.* p. 233.
 CAMERON, A. F., FERGUSON, G. & MORRIS, D. G. (1968). *Chem. Comm.* 6, 316.
 CRUICKSHANK, D. W. J. & ROBERTSON, A. P. (1953). *Acta Cryst.* 6, 698.

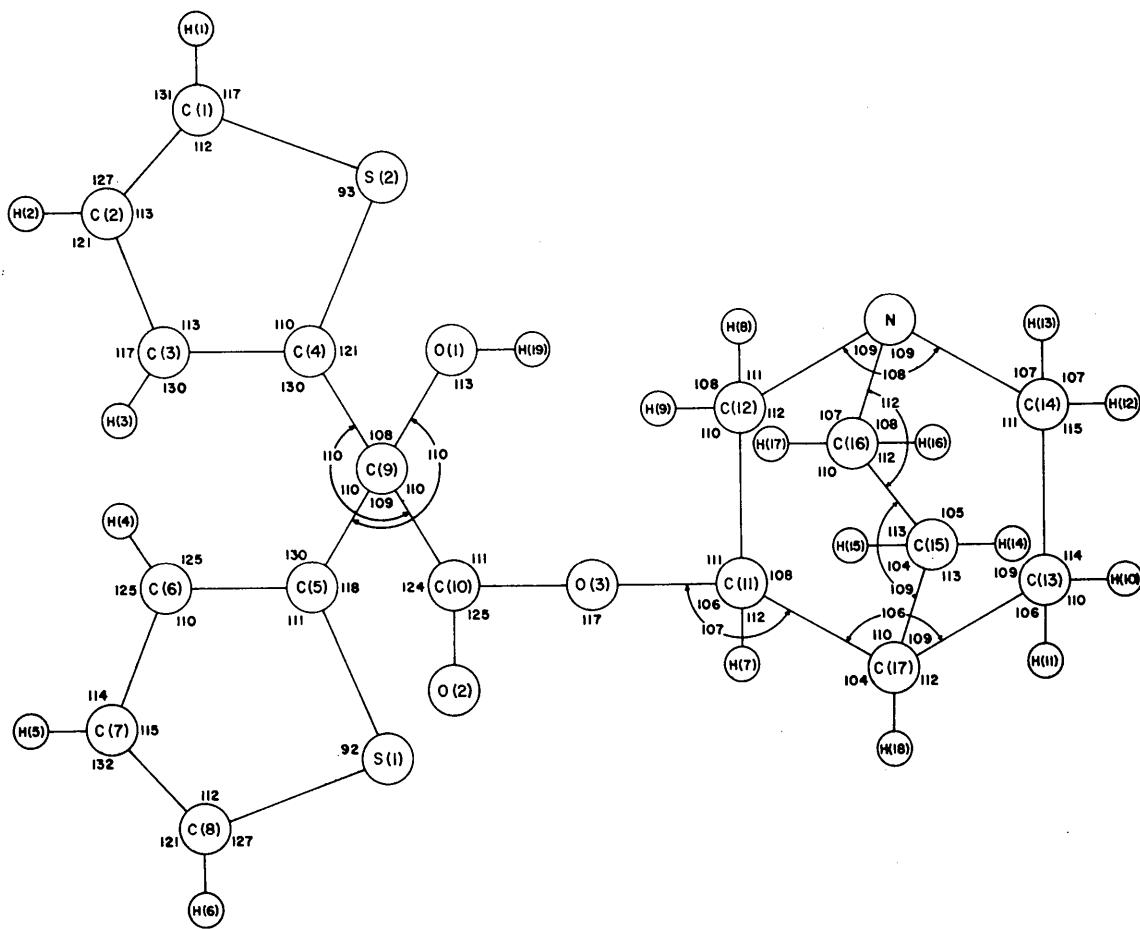


Fig. 4. Bond angles (°).

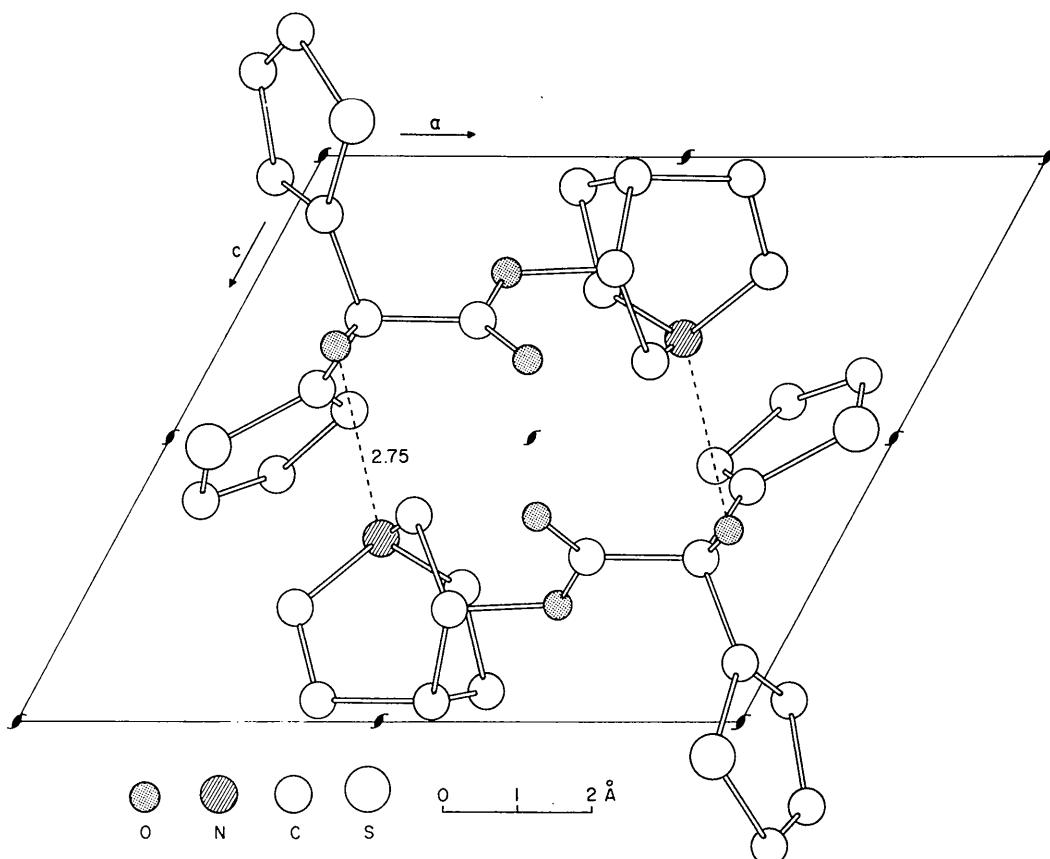


Fig. 5. Quinuclidinyl di- α,α' -thienylglycolate. Projection of the structure along the b axis showing intermolecular hydrogen bonds (\AA). All hydrogen atoms and adjacent symmetry related molecules are omitted.

- DONOHUE, J. (1950). *J. Amer. Chem. Soc.* **72**, 2701.
 GABEL, N. W. & ABOOD, L. G. (1965). *J. Med. Chem.* **8**, 616.
 HUDSON, P. & ROBERTSON, J. H. (1962). *Acta Cryst.* **15**, 913.
 HUDSON, P. & ROBERTSON, J. H. (1964). *Acta Cryst.* **17**, 1497
International Tables for X-ray Crystallography (1962). Vol. III. Birmingham: Kynoch Press.
 JOHNSON, C. K. (1965). *ORTEP, A Fortran Thermal Ellipsoid Plot Program for Crystal Structure Illustrations*. ORNL-3794, Oak Ridge National Laboratory, Tennessee.
 MEYERHÖFFER, A. & CARLSTRÖM, D. (1969). *Acta Cryst.* **B25**, 1119.
 MOMANY, F. A. (1964). *Dissert. Abstr.* **24**, 2714.
 NARDELLI, M., FAVA, G. & GIRALDI, G. (1962). *Acta Cryst.* **15**, 737.
 NYBERG, K., ÖSTMAN, B. & WALLERBERG, G. (1970). *Acta Chem. Scand.* In the press.
 PAULING, L. (1960). *The Nature of the Chemical Bond*, p. 189. Ithaca: Cornell Univ. Press.
 RINGERTZ, H. (1969). Thesis, Stockholm.
 STEWART, J. M. & HIGH, D. (1965). *X-ray 63, Program System for X-ray Crystallography* for the IBM 7090. Univ. of Washington, Seattle and Univ. of Maryland, College Park.
 VISSER, G. J., HEERES, G. J., WOLTERS, J. & VOS, A. (1968). *Acta Cryst.* **B24**, 467.
 WALKER, L. A., FOLTING, K. & MERRITT, L. L. JR (1969). *Acta Cryst.* **B25**, 88.