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## Investigations of Dithienylglycollic Esters. II. The Crystal and Molecular Structure of Quinuclidinyl Di- $\alpha, \alpha'$ -thienylglycollate

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Crystals of quinuclidinyl di- $\alpha, \alpha'$ -thienylglycollate are monoclinic, space group  $P2_1$  with two formula units in a cell having the dimensions  $a = 9.762$ ,  $b = 11.348$ ,  $c = 8.556$  Å,  $\beta = 118.405^\circ$ . The crystal structure was solved by a three-dimensional Patterson synthesis and refined by three-dimensional least-squares to a final  $R$  value of 0.042. The two thiophene rings make an angle of  $96^\circ$  to each other. The glycollic bridge connecting the thiophene rings with the quinuclidine is almost planar. The quinuclidine skeleton is twisted  $5^\circ$  around its threefold axis. An intermolecular hydrogen bond is present between the hydroxyl hydrogen atom of the glycollic bridge and the nitrogen atom in the quinuclidine.

### Introduction

The crystal structure of the anticholinergic drug, quinuclidinyl di- $\alpha, \alpha'$ -thienylglycollate,



has been solved as part of our programme on the correlation between molecular configuration and pharmacological action of psychotomimetic compounds. The crystal structure of the closely related quinuclidinyl benzilate has been determined in the hydrobromide salt (Meyerhöffer & Carlström, 1969). As one of the important factors in psychotomimetic drugs is probably the availability of a free electron pair on a nitrogen atom (Gabel & Abood, 1965), it was regarded as important to solve the structure of a psychoactive base. All attempts to determine the structure of the base of the benzilate proved unsuccessful. For this reason the equally active thiophene compound was chosen as it could be analyzed directly by the 'heavy-atom' method. Another advantage was that the hydrogen positions could be determined, which was not possible in the case of the hydrobromide salt because of the strong interference of the heavy bromine atoms.

### Experimental

The synthesis of quinuclidinyl di- $\alpha, \alpha'$ -thienylglycollate has been described by Nyberg, Östman & Wallerberg

(1970). Large, colourless, optically perfect crystals were prepared through recrystallization in acetone during slow cooling. The unit-cell dimensions were determined from Weissenberg photographs and from diffractometer measurements. The density of single crystals was established by flotation in mixed solvents.

### Crystal data

Quinuclidinyl di- $\alpha, \alpha'$ -thienylglycollate,  $C_{17}H_{19}NO_3S_2$   
F. W. 349.47

$$\begin{aligned} a &= 9.762 \pm 0.005 \text{ \AA} \\ b &= 11.348 \pm 0.005 \text{ \AA} \\ c &= 8.556 \pm 0.005 \text{ \AA} \\ \beta &= 118.405^\circ \pm 0.010^\circ \\ V &= 833.72 \text{ \AA}^3 \\ D_m &= 1.393 \pm 0.005 \text{ g.cm}^{-3} \\ D_x (Z=2) &= 1.392 \pm 0.002 \text{ g.cm}^{-3} \end{aligned}$$

The only systematic absences were  $0k0$  when  $k$  is odd. Since the molecule is asymmetric and the unit cell contains two molecules the space group can only be  $P2_1$ .

Intensity data were collected with a Philips automatic single-crystal diffractometer (PAILRED) using monochromatic  $Mo K\alpha$  radiation. The crystal, having the dimensions  $0.42 \times 0.62 \times 0.41$  mm, was oriented about the  $b$  axis and 15 layers corresponding to the entire Cu sphere were recorded. There were 2171 independent reflexions of which 146 were too weak to be

observed. The unobserved reflexions were given half the intensity of the threshold value. The intensities were corrected for corresponding background,  $\alpha_1$ - $\alpha_2$  separation, Lorentz and polarization factors, but not for extinction or absorption. The structure amplitudes were placed approximately on an absolute scale by Wilson statistics. The above calculations were performed with the program system of Ringertz (1969). The X-ray 63 program system (Stewart & High, 1965) was used during the following procedures except for the perspective drawing, which was computed with the *ORTEP* program (Johnson, 1965). All calculations were carried out on an IBM 7090 computer.

### Determination of the structure

A three-dimensional Patterson function was calculated from the corrected set of diffractometer data. In the Harker section  $y = \frac{1}{2}$ , two heavy peaks were located. All four combinations of these two peaks were tried as S-S rotation peaks, but calculations of structure amplitudes showed essentially no differences in the reliability indices  $R$  ( $R = \sum |F_o - F_c| / \sum |F_c|$ ) between the different alternatives. Instead a combination of one of

the heavy peaks with another peak having roughly half the intensity of the first chosen caused a marked drop in the  $R$  value, indicating that the true S positions had been found. The electron density map, based upon the sulphur signs only, showed, together with the sulphur maxima, additional peaks in the vicinity of these maxima and one of the five-membered rings could be identified. Iterated Fourier refinement gave the other five-membered ring and finally the entire molecule. The  $R$  value at this stage was 0.19 with an overall temperature factor  $B = 3.5 \text{ \AA}^2$ . It was now possible to interpret correctly the Patterson function; one of the two heavy peaks in the Harker section  $y = \frac{1}{2}$  was found to be composed of three coincident vectors, namely O-O, S-O and S-C.

### Refinement procedure

The initial refinement was carried out by a block-diagonal three-dimensional least-squares program because of the large number of parameters involved. However, for the final refinement a full-matrix program was used. In this case the refinement had to be carried out in two steps. Initially only the strongest 1650 of the 2171 reflexions were used and the structure amplitudes were

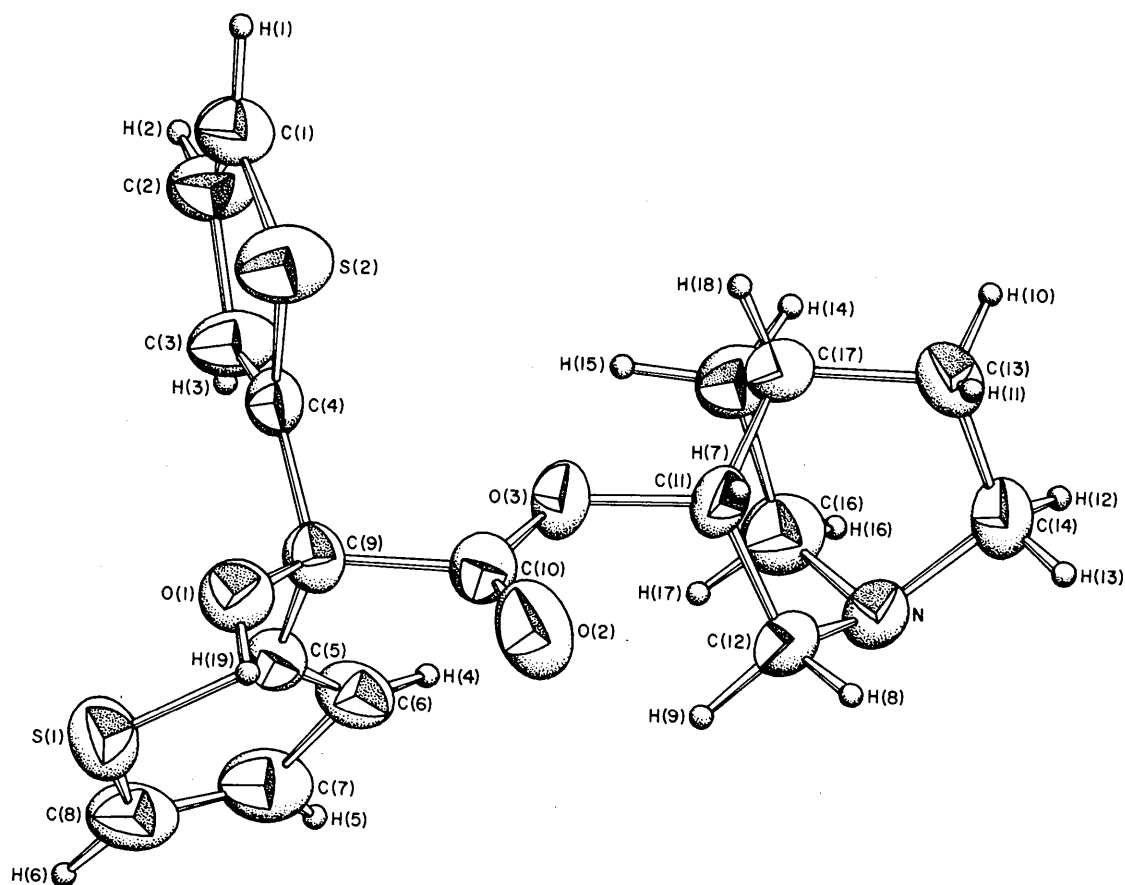


Fig. 1. A perspective drawing of the quinuclidinyl di- $\alpha, \alpha'$ -thienylglycollate molecule, showing the ellipsoids of thermal motion with a probability of 50%.



INVESTIGATIONS OF DITHIONOGLYCOLLIC ESTERS. II

Table 1 (cont.)

Table with multiple columns (H, K, L, FO, FC, AC, BC) and rows of numerical data. The table is divided into four main sections by column headers.

Table 1 (cont.)

Table with 48 columns (H, K, L, FL, FC, AC, BC) and multiple rows of numerical data. The table is organized into 12 groups of 4 columns each, with each group containing 12 rows of data.

given unit weight. The atomic scattering factors were those listed in *International Tables for X-ray Crystallography* (1962).

The first two cycles of block-diagonal refinement with non-hydrogen atoms only and isotropic temperature factors yielded an  $R$  value of 0.10. When anisotropic temperature factors were introduced the  $R$  value decreased to 0.075. A Fourier difference synthesis prepared at this stage made it possible to locate all of the 19 hydrogen atoms. With the hydrogen atoms included and given isotropic temperature factors, three cycles of refinement of the non-hydrogen atoms lowered the  $R$  value to 0.036. When also the positional parameters and the individual isotropic temperature factors of the hydrogen atoms were refined for one cycle the  $R$  value decreased to 0.033. The standard deviations of the atomic parameters were now quite satisfactory. In the next step the structure amplitudes were weighted. The weighting scheme adopted was  $w = (a + bF_o + cF_o^2)^{-1}$  where  $a$ ,  $b$  and  $c$  were derived from a quadratic curve of best fit for  $|F_o - F_c|$  as a function of  $F_o$ . The numerical values of  $a$ ,  $b$  and  $c$  were  $a = 0.12$ ,  $b = 0.017$ ,  $c = -0.0004$  for  $F_o \geq 10$  and  $a = 0.282$ ,  $b = c = 0$  for  $F_o < 10$ . A final refinement including all of the 2171 weighted structure factors resulted in  $R = 0.042$ , *i.e.* somewhat higher than the lowest value for the 1650 reflexions, but the standard deviations of the atomic parameters were further decreased. The average of the shift to error ratio of the parameters was 0.20 for the non-hydrogen atoms and 0.85 for the hydrogen atoms. Observed and calculated structure factors from this last cycle of refinement are listed in Table 1. The positional and thermal parameters for the final structure are given in Tables 2 and 3 respectively.

Table 2. *Final fractional atomic coordinates and their standard deviations (in parentheses)*

	$x/a$	$y/b$	$z/c$
S(1)	0.0594 (1)	0.3500	1.0177 (1)
S(2)	0.0114 (1)	0.4893 (1)	0.4370 (1)
O(1)	0.1598 (3)	0.5242 (2)	0.8389 (3)
O(2)	0.4397 (4)	0.4686 (4)	0.8619 (5)
O(3)	0.3416 (3)	0.3041 (3)	0.7041 (3)
N	0.6342 (3)	0.0976 (3)	0.8229 (4)
C(1)	-0.1344 (5)	0.4053 (4)	0.2803 (5)
C(2)	-0.1562 (5)	0.3076 (4)	0.3496 (5)
C(3)	-0.0536 (4)	0.2979 (4)	0.5352 (4)
C(4)	-0.0466 (4)	0.3899 (3)	0.6024 (4)
C(5)	0.1678 (4)	0.3181 (3)	0.9136 (4)
C(6)	0.2262 (5)	0.2028 (4)	0.9513 (5)
C(7)	0.1756 (6)	0.1463 (5)	1.0649 (5)
C(8)	0.0876 (5)	0.2132 (5)	1.1094 (5)
C(9)	0.1767 (4)	0.4104 (3)	0.7889 (4)
C(10)	0.3366 (4)	0.3995 (4)	0.7912 (4)
C(11)	0.4863 (4)	0.2835 (4)	0.6974 (5)
C(12)	0.6058 (4)	0.2171 (4)	0.8642 (5)
C(13)	0.6004 (5)	0.1804 (5)	0.5370 (5)
C(14)	0.7027 (5)	0.1058 (5)	0.7013 (6)
C(15)	0.3748 (5)	0.0912 (4)	0.5515 (5)
C(16)	0.4851 (5)	0.0345 (4)	0.7332 (6)
C(17)	0.4461 (4)	0.2060 (4)	0.5357 (5)
H(1)	-0.176 (6)	0.430 (5)	0.157 (7)
H(2)	-0.238 (7)	0.244 (6)	0.284 (8)

Table 2 (cont.)

H(3)	-0.058 (5)	0.239 (5)	0.582 (7)
H(4)	0.275 (4)	0.178 (4)	0.923 (5)
H(5)	0.215 (7)	0.069 (6)	1.099 (8)
H(6)	0.046 (6)	0.183 (5)	1.183 (7)
H(7)	0.522 (5)	0.360 (4)	0.684 (5)
H(8)	0.709 (4)	0.266 (4)	0.922 (5)
H(9)	0.565 (5)	0.210 (5)	0.956 (6)
H(10)	0.582 (5)	0.147 (5)	0.434 (6)
H(11)	0.651 (6)	0.256 (5)	0.549 (7)
H(12)	0.718 (6)	0.017 (5)	0.674 (7)
H(13)	0.802 (6)	0.138 (5)	0.765 (7)
H(14)	0.367 (7)	0.030 (5)	0.459 (8)
H(15)	0.268 (5)	0.117 (4)	0.541 (6)
H(16)	0.506 (7)	-0.043 (6)	0.722 (8)
H(17)	0.440 (5)	0.037 (4)	0.809 (6)
H(18)	0.379 (5)	0.254 (4)	0.438 (5)
H(19)	0.217 (6)	0.538 (5)	0.932 (6)

## Description and discussion of the structure

### The geometry of the molecule

The atomic nomenclature and the configuration of the molecule is illustrated in Fig. 1 which also shows the ellipsoids of thermal motion scaled to include 50 per cent probability. As can be seen, the molecule is composed of two planar  $\alpha$ -substituted thiophene rings connected to a quinuclidine nucleus by means of an almost planar bridge, called the glycollic bridge in the following. The best planes through each of the two five-membered rings were calculated with and without the sulphur atoms included. In the former case all the atoms deviated significantly from the planes whereas only the sulphur atoms deviated significantly in the latter case. These deviations amounted to 0.010 Å for S(1) and 0.014 Å for S(2). This is slightly less than found in thenoic acid. Thus, Nardelli, Fava & Giraldi (1962) observed that the sulphur atom was displaced 0.028 Å from a plane through the carbon atoms in the ring of  $\alpha$ -thenoic acid. Later, Hudson & Robertson (1962, 1964) confirmed this result. They found the

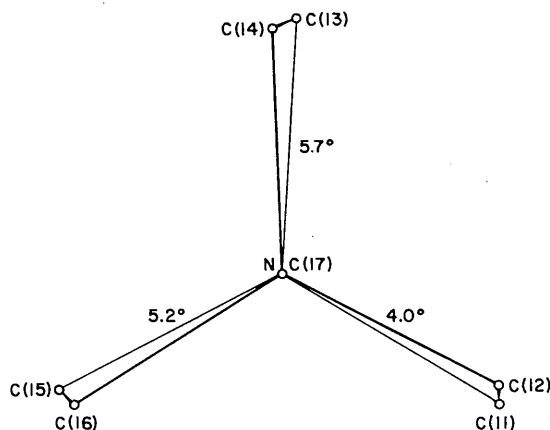


Fig. 2. The quinuclidine part of the molecule seen along its threefold axis, showing the angle of twist between the upper and lower halves.

deviation 0.02 Å of the sulphur atoms in  $\alpha$ - and  $\beta$ -thenoic acids. The equations of the planes and the deviations of the atoms in the present structure are given in Table 4. The planes of the rings make an angle of 96.4° to each other.

The glycollic bridge connecting the thiophene rings to the quinuclidine part of the molecule is perfectly planar except for the hydroxyl oxygen atom, O(1), which deviates 0.34 Å from a plane defined by the atoms O(2), O(3), C(9), C(10), C(11). The equation of this plane and the atomic deviations are given in Table 4. The angles between the plane of the glycollic bridge and the thiophene rings C(1) to C(4) and C(5) to C(8) are 93° and 119° respectively.

The quinuclidine cage is almost symmetric in shape. However, when seen along its axis of symmetry,

C(17)...N, it exhibits a slight distortion illustrated in Fig. 2. The mean angle of twist is  $5.0 \pm 0.7^\circ$  which is the same as that found in our earlier investigation of the quinuclidinyl benzilate (Meyerhöffer & Carlström, 1969). The question regarding the symmetry of quinuclidine and similar cage-like structures has already been discussed in this work. Recently, a new structure determination of a bicyclo[2,2,2]octane derivative has been reported by Cameron, Ferguson & Morris (1968) who likewise observed an angle of twist ( $3^\circ$ ) about the threefold axis of the carbocycle. According to these authors the distortion is adopted solely to release strain from torsional and nonbonded intramolecular interactions.

The overall geometry of the present molecule is very similar to that of the quinuclidinyl benzilate in spite

Table 3. *Anisotropic parameters and standard deviations (in parentheses)*

$\beta_{ij}$  are the coefficients in the expression  $\exp[-(\beta_{11}h^2 + \beta_{22}k^2 + \beta_{33}l^2 + 2\beta_{12}hk + 2\beta_{13}hl + 2\beta_{23}kl)]$ .

	$\beta_{11}$	$\beta_{22}$	$\beta_{33}$	$\beta_{12}$	$\beta_{13}$	$\beta_{23}$
S(1)	0.0140 (1)	0.0099 (1)	0.0208 (2)	0.0004 (1)	0.0111 (1)	0.0007 (1)
S(2)	0.0186 (2)	0.0080 (1)	0.0183 (2)	-0.0019 (1)	0.0070 (2)	0.0030 (1)
O(1)	0.0130 (4)	0.0057 (2)	0.0159 (4)	0.0002 (2)	0.0057 (3)	-0.0023 (2)
O(2)	0.0153 (5)	0.0124 (4)	0.0289 (7)	-0.0066 (3)	0.0134 (5)	-0.0091 (5)
O(3)	0.0086 (3)	0.0064 (2)	0.0191 (5)	0.0004 (2)	0.0066 (3)	-0.0009 (3)
N	0.0098 (4)	0.0082 (3)	0.0144 (5)	0.0016 (3)	0.0051 (4)	0.0010 (3)
C(1)	0.0150 (6)	0.0100 (3)	0.0133 (6)	0.0013 (4)	0.0061 (5)	0.0007 (4)
C(2)	0.0146 (6)	0.0101 (4)	0.0135 (6)	-0.0025 (4)	0.0046 (5)	-0.0016 (4)
C(3)	0.0142 (5)	0.0068 (3)	0.0130 (5)	-0.0026 (3)	0.0037 (4)	0.0002 (3)
C(4)	0.0094 (4)	0.0055 (2)	0.0137 (5)	0.0008 (2)	0.0066 (4)	0.0002 (3)
C(5)	0.0093 (4)	0.0066 (3)	0.0112 (5)	-0.0005 (3)	0.0051 (4)	-0.0007 (3)
C(6)	0.0157 (6)	0.0067 (3)	0.0133 (6)	0.0006 (3)	0.0081 (5)	0.0014 (3)
C(7)	0.0185 (7)	0.0084 (4)	0.0156 (7)	-0.0009 (4)	0.0072 (6)	0.0019 (4)
C(8)	0.0152 (6)	0.0114 (5)	0.0132 (6)	-0.0039 (5)	0.0061 (5)	0.0010 (4)
C(9)	0.0102 (4)	0.0049 (2)	0.0143 (5)	0.0003 (3)	0.0065 (4)	-0.0007 (3)
C(10)	0.0097 (4)	0.0070 (3)	0.0144 (5)	-0.0010 (3)	0.0058 (4)	-0.0004 (3)
C(11)	0.0072 (4)	0.0075 (3)	0.0157 (6)	0.0005 (3)	0.0049 (4)	0.0005 (3)
C(12)	0.0085 (4)	0.0090 (3)	0.0132 (5)	0.0015 (3)	0.0032 (4)	-0.0001 (4)
C(13)	0.0144 (6)	0.0126 (5)	0.0167 (7)	0.0025 (4)	0.0102 (6)	0.0014 (5)
C(14)	0.0123 (5)	0.0120 (5)	0.0193 (7)	0.0036 (4)	0.0089 (5)	0.0011 (5)
C(15)	0.0119 (5)	0.0093 (4)	0.0162 (7)	-0.0010 (4)	0.0044 (5)	-0.0026 (4)
C(16)	0.0142 (6)	0.0086 (4)	0.0198 (8)	-0.0014 (4)	0.0072 (6)	0.0002 (5)
C(17)	0.0106 (5)	0.0099 (4)	0.0122 (5)	0.0019 (3)	0.0045 (4)	0.0010 (4)

Table 4. *Equations of the best planes*

The equations are of the form  $m_1X + m_2Y + m_3Z = d$ ,  $X, Y, Z$  being the orthogonal coordinates in Å.

	Plane	$m_1$	$m_2$	$m_3$	$d$
<i>A</i>	C(1), C(2), C(3), C(4), S(2)	0.8587	-0.4743	-0.1941	-4.6989
<i>B</i>	C(1), C(2), C(3), C(4)	0.8563	-0.4787	-0.1939	-4.7087
<i>C</i>	C(5), C(6), C(7), C(8), S(1)	0.4995	0.3280	0.8018	5.6614
<i>D</i>	C(5), C(6), C(7), C(8)	0.4975	0.3308	0.8019	5.6733
<i>E</i>	C(9), C(10), C(11), O(2), O(3)	-0.0523	0.5289	0.8471	2.6423

Deviations (Å) from the best planes

	Plane <i>A</i>	Plane <i>B</i>	Plane <i>C</i>	Plane <i>D</i>	Plane <i>E</i>
C(1)	-0.002	C(1) 0.002	C(5) 0.004	C(5) 0.001	C(9) -0.002
C(2)	-0.002	C(2) -0.004	C(6) -0.003	C(6) -0.001	C(10) -0.000
C(3)	0.006	C(3) 0.004	C(7) 0.000	C(7) 0.001	C(11) -0.002
C(4)	-0.006	C(4) -0.002	C(8) 0.002	C(8) -0.001	O(2) 0.001
S(2)	0.004		S(1) -0.003		O(3) 0.003
		S(2) 0.014		S(1) -0.010	
					O(1) 0.344
					H(19) -0.152

of the latter being a salt. The main differences are due to twists around the C(4)–C(9) and C(5)–C(9) bonds connecting the rings to the glycollic bridge, a 23° twist around the C(9)–C(10) bond and a 20° twist around the O(3)–C(11) bond. A closer comparison of the two structures seems to be of little value because of the much lower accuracy in the determination of the benzilate structure.

#### Bond lengths and bond angles

The bond lengths and bond angles with estimated standard deviations are listed in Tables 5 and 6 and illustrated in Fig. 3 and 4. No correction has been made for thermal motion. The thiophene rings were non-symmetric when tested with the criteria of significance proposed by Cruickshank & Robertson (1953). Thus the C(1)–C(2) bond is significantly shorter than the C(3)–C(4) bond and the same is true for corresponding bonds in the other thiophene ring. C(1)–C(2) (1.321 Å) and C(7)–C(8) (1.331 Å) are moreover slightly shorter than the standard double bond length, of 1.337 Å. Such non-symmetric thiophene rings with short C–C bonds have earlier been described by Block, Filippakis & Schmidt (1967) in  $\alpha$ -thienylacrylic acid.

Table 5. *Interatomic distances and standard deviations (in parentheses)*

S(2)–C(4)	1.712 (4) Å	C(11)–C(12)	1.543 (5) Å
S(2)–C(1)	1.708 (4)	C(13)–C(14)	1.535 (6)
C(1)–C(2)	1.321 (7)	C(15)–C(16)	1.548 (6)
C(2)–C(3)	1.422 (5)	C(11)–C(17)	1.524 (6)
C(3)–C(4)	1.358 (5)	C(13)–C(17)	1.529 (7)
		C(15)–C(17)	1.514 (7)
S(1)–C(5)	1.714 (5)	C(12)–N	1.461 (6)
S(1)–C(8)	1.702 (5)	C(14)–N	1.483 (7)
C(8)–C(7)	1.331 (8)	C(16)–N	1.469 (6)
C(7)–C(6)	1.434 (8)		
C(6)–C(5)	1.402 (6)	C(11)–H(7)	0.96 (5)
		C(12)–H(8)	1.04 (4)
C(1)–H(1)	0.98 (6)	C(12)–H(9)	1.04 (6)
C(2)–H(2)	1.02 (6)	C(13)–H(10)	0.89 (6)
C(3)–H(3)	0.79 (6)	C(13)–H(11)	0.96 (6)
C(6)–H(4)	0.69 (5)	C(14)–H(12)	1.06 (6)
C(7)–H(5)	0.94 (6)	C(14)–H(13)	0.93 (5)
C(8)–H(6)	0.96 (7)	C(15)–H(14)	1.03 (7)
		C(15)–H(15)	1.04 (5)
C(4)–C(9)	1.511 (4)	C(16)–H(16)	0.91 (7)
C(5)–C(9)	1.527 (6)	C(16)–H(17)	0.94 (6)
C(9)–C(10)	1.556 (6)	C(17)–H(18)	0.95 (4)
C(9)–O(1)	1.395 (5)		
C(10)–O(2)	1.190 (5)	O(1)····O(2)	2.719 (5)
C(10)–O(3)	1.329 (5)	O(1)····N <sup>+</sup>	2.752 (4)
O(3)–C(11)	1.460 (5)	H(19)····O(2)	2.63 (6)
		H(19)····N <sup>+</sup>	2.02 (4)
O(1)–H(19)	0.74 (4)		

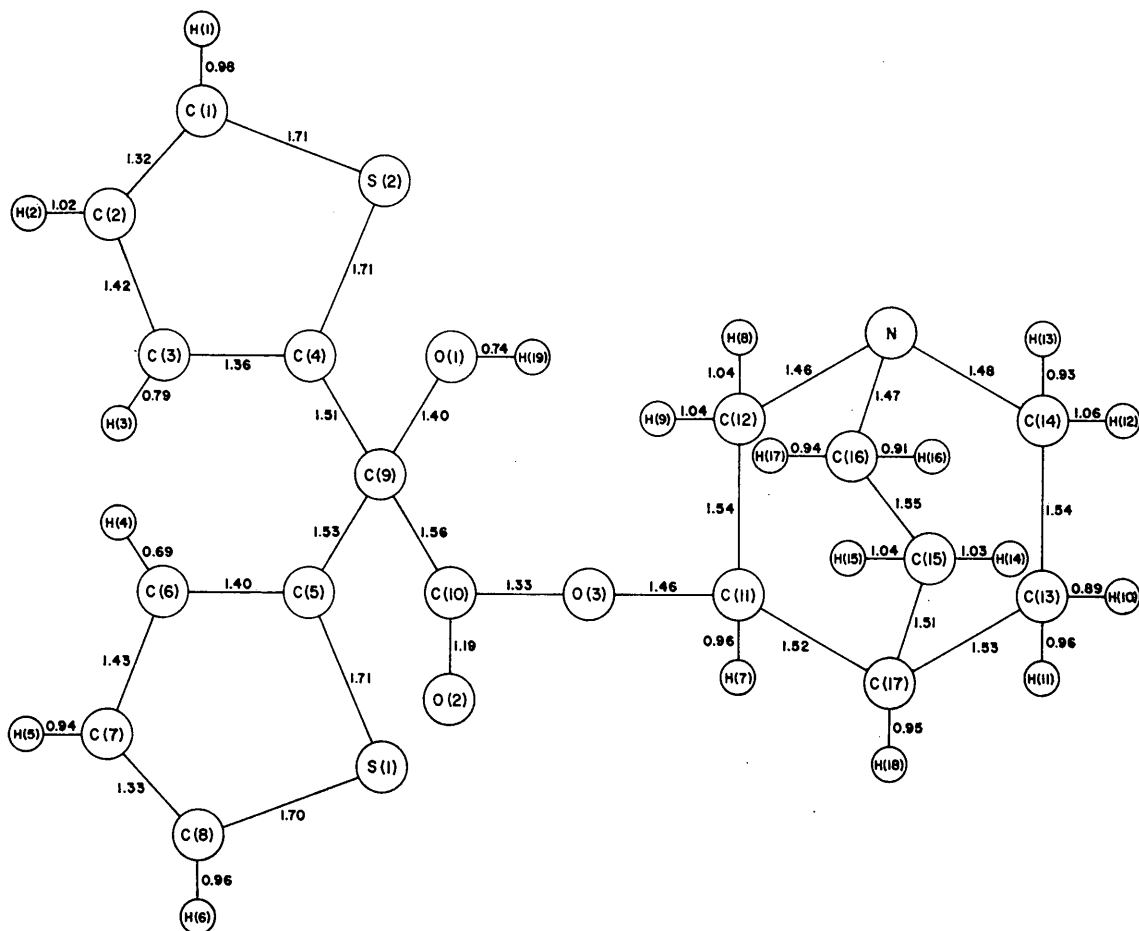


Fig. 3. Bond lengths (Å).



On comparing corresponding bond lengths in the two thiophene rings all bonds were found to be identical except for the C(3)–C(4) and C(5)–C(6) which were significantly different. Averaging the bond lengths of the thiophene rings, the mean S–C distance, 1.709 Å, is almost midway between a pure single bond, 1.815 Å and a pure double bond, 1.556 Å, as given by Bak, Christensen, Hansen-Nygaard & Rastrup-Andersen, (1961). The mean C–C bond length, 1.428 Å, opposite the sulphur atom and the mean C–C bond length, 1.353 Å, adjacent to the C–S bond deviate symmetrically with respect to the C–C bond length, 1.395 Å, of benzene. The mean bond lengths and the bond angles are all in excellent agreement with the earlier determinations of a  $\alpha$ -substituted thiophene (Nardelli, Fava & Giraldi, 1962) as well as calculations of the thiophene structure from micro wave spectra (Bak, Christensen, Hansen-Nygaard & Rastrup-Andersen, 1961) and from electron diffraction (Momany, 1964). Large varia-

tions in bond lengths are reported for a  $\beta$ -substituted thiophene ring (Hudson & Robertson, 1964). On the other hand Visser, Heeres, Wolters & Vos (1968) have found normal thiophene bond lengths in  $\alpha, \alpha'$ - and  $\beta, \beta'$ -thienyls within the limits of the rather large experimental errors. In the rest of the molecule all bond lengths and bond angles are close to their normal values. The mean paraffinic C–C distance, 1.532 Å, the mean C–N distance, 1.471 Å and the mean C–H distance, 0.95 Å are all quite acceptable and so are the C=O and C–O bond lengths.

#### Hydrogen bonds and packing of the molecules

The observed planarity of the connecting glycollic bridge together with the short O(1) . . . O(2) distance, 2.719 Å, at first indicated the presence of an intramolecular hydrogen bond as presumed for the quinuclidinyl benzilate molecule (Meyerhöffer & Carlström, 1969). However, the short O(1) . . . N'(1-x, y+½, 1-z)

Table 6. *Interatomic angles and standard deviations (in parentheses)*

C(4)–S(2)–C(1)	92.6 (2)°	C(5)–S(1)–C(8)	92.4 (2)°
S(2)–C(1)–C(2)	111.6 (3)	S(1)–C(8)–C(7)	111.9 (4)
C(1)–C(2)–C(3)	112.8 (4)	C(8)–C(7)–C(6)	114.6 (5)
C(2)–C(3)–C(4)	113.3 (4)	C(7)–C(6)–C(5)	109.9 (4)
C(3)–C(4)–S(2)	109.6 (2)	C(6)–C(5)–S(1)	111.2 (3)
C(3)–C(4)–C(9)	129.8 (3)	C(6)–C(5)–C(9)	130.4 (4)
S(2)–C(4)–C(9)	120.6 (3)	S(1)–C(5)–C(9)	118.2 (3)
S(2)–C(1)–H(1)	117 (3)	S(1)–C(8)–H(6)	127 (3)
C(2)–C(1)–H(1)	131 (3)	C(7)–C(8)–H(6)	121 (3)
C(1)–C(2)–H(2)	127 (4)	C(8)–C(7)–H(5)	132 (5)
C(3)–C(2)–H(2)	121 (4)	C(6)–C(7)–H(5)	114 (5)
C(2)–C(3)–H(3)	117 (3)	C(7)–C(6)–H(4)	125 (4)
C(4)–C(3)–H(3)	130 (3)	C(5)–C(6)–H(4)	125 (4)
C(4)–C(9)–C(5)	109.5 (3)	C(11)–C(12)–N	112.1 (3)
C(4)–C(9)–C(10)	109.6 (3)	C(13)–C(14)–N	111.3 (4)
C(5)–C(9)–C(10)	109.1 (3)	C(15)–C(16)–N	111.5 (4)
C(4)–C(9)–O(1)	107.7 (3)	C(11)–C(17)–C(13)	106.0 (3)
C(5)–C(9)–O(1)	111.5 (3)	C(13)–C(17)–C(15)	109.3 (4)
C(10)–C(9)–O(1)	109.5 (3)	C(15)–C(17)–C(11)	110.3 (4)
C(9)–C(10)–O(2)	123.7 (4)	C(12)–N–C(14)	108.2 (4)
C(9)–C(10)–O(3)	111.2 (3)	C(14)–N–C(16)	109.2 (3)
O(2)–C(10)–O(3)	125.1 (4)	C(16)–N–C(12)	108.9 (3)
C(10)–O(3)–C(11)	116.6 (3)	C(9)–O(1)–H(19)	113 (4)
O(3)–C(11)–C(12)	111.2 (4)	O(3)–C(11)–H(7)	106 (3)
O(3)–C(11)–C(17)	107.0 (3)	C(12)–C(11)–H(7)	113 (2)
C(12)–C(11)–C(17)	108.3 (3)	C(17)–C(11)–H(7)	112 (3)
C(14)–C(13)–C(17)	108.7 (4)	C(11)–C(12)–H(8)	108 (2)
C(16)–C(15)–C(17)	108.5 (3)	C(17)–C(15)–H(14)	113 (4)
N—C(12)–H(8)	111 (2)	C(16)–C(15)–H(15)	113 (3)
C(11)–C(12)–H(9)	110 (3)	C(17)–C(15)–H(15)	104 (3)
N—C(12)–H(9)	107 (3)	H(14)–C(15)–H(15)	114 (4)
H(8)–C(12)–H(9)	108 (4)	C(15)–C(16)–H(16)	112 (4)
C(14)–C(13)–H(10)	114 (3)	N—C(16)–H(16)	108 (4)
C(17)–C(13)–H(10)	110 (3)	C(15)–C(16)–H(17)	110 (2)
C(14)–C(13)–H(11)	108 (3)	N—C(16)–H(17)	107 (2)
C(17)–C(13)–H(11)	106 (4)	H(16)–C(16)–H(17)	108 (6)
H(10)–C(13)–H(11)	110 (6)	C(11)–C(17)–H(18)	104 (3)
C(13)–C(14)–H(12)	115 (3)	C(13)–C(17)–H(18)	112 (3)
N—C(14)–H(12)	105 (4)	C(15)–C(17)–H(18)	115 (3)
C(13)–C(14)–H(13)	111 (4)	O(1)–H(19) . . . O(2)	89 (5)
N—C(14)–H(13)	107 (4)	O(1)–H(19) . . . N'	173 (5)
H(12)–C(14)–H(13)	107 (5)		
C(16)–C(15)–H(14)	105 (3)		

distance, 2.752 Å, seemed equally well suited to an intermolecular hydrogen bond. When the position of H(19) was established it became evident that the H(19)...O(2) distance, 2.63 Å, was much too long to permit such a bond whereas the H(19)...N' distance, 2.02 Å was quite favourable. The existence of a hydrogen bond between O(1)H(19)...N' is further supported by the fact that the hydrogen atom lies almost on the straight line between the oxygen atom and the nitrogen atom, the O(1)–H(19)...N' angle being 173°. The hydrogen bond system is illustrated in Fig. 5.

The oxygen atom O(1) of the hydroxyl group is facing the sulphur atoms of the thiophene rings, the distances being as short as 2.939 Å and 3.056 Å. This verifies earlier findings that the van der Waals radius of sulphur is closer to 1.60 Å (Donohue, 1950; Nardelli, Fava & Giraldi, 1962), than to the value of 1.85 Å given by Pauling (1960) and even to the value of 1.72 Å supposed by Walker, Folting & Merritt (1969). All other distances are in accordance with ordinary requirements.

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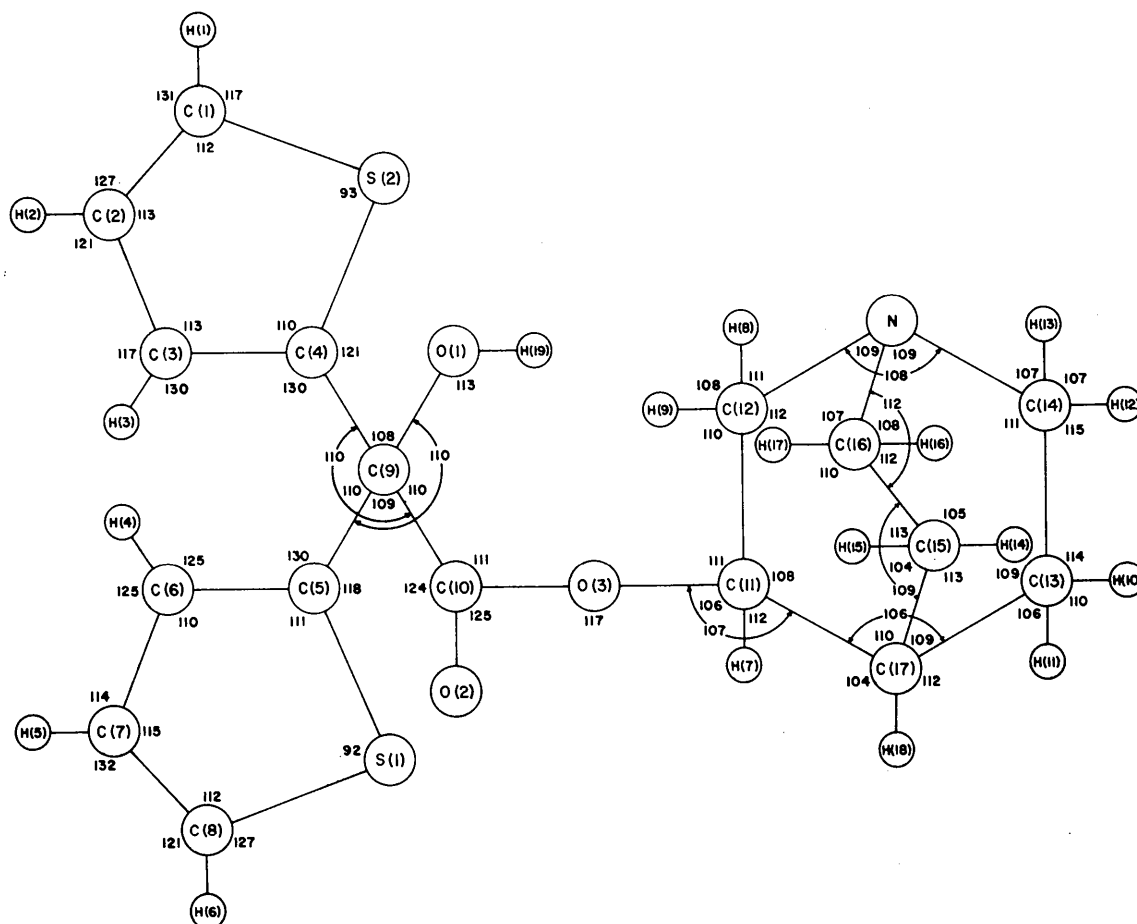


Fig. 4. Bond angles (°).

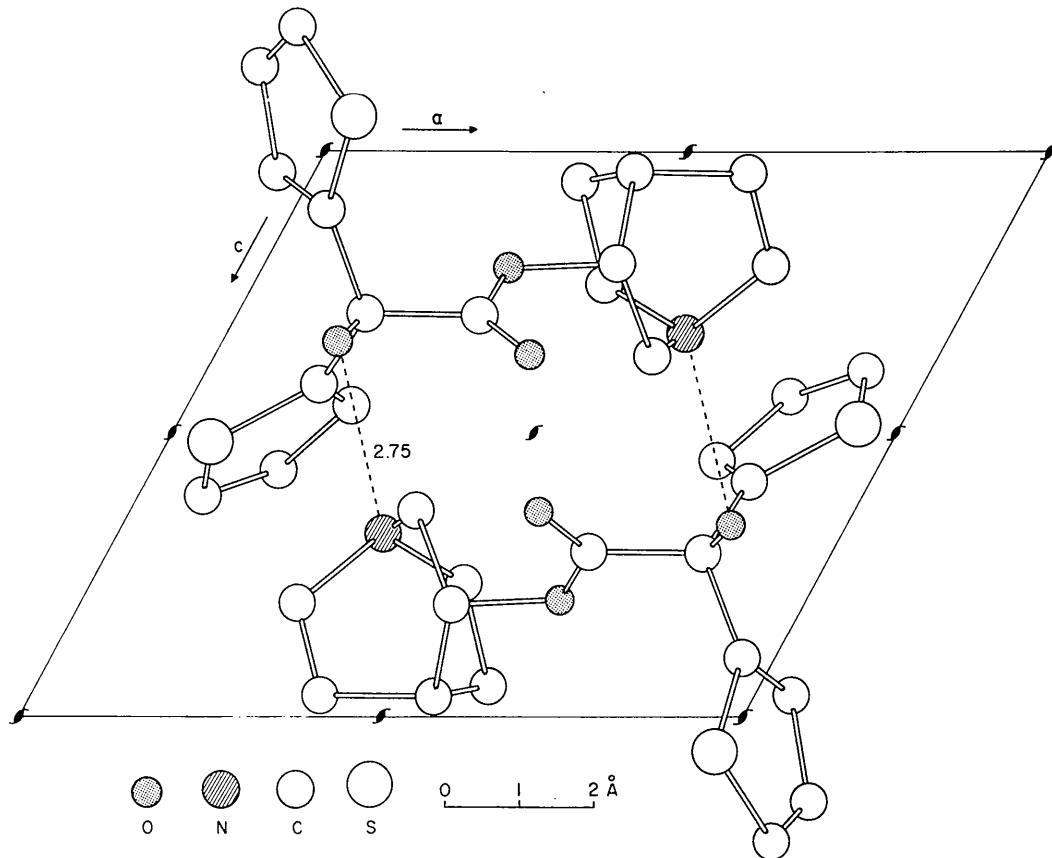


Fig. 5. Quinuclidinyl di- $\alpha,\alpha'$ -thienylglycollate. Projection of the structure along the  $b$  axis showing intermolecular hydrogen bonds ( $\text{\AA}$ ). All hydrogen atoms and adjacent symmetry related molecules are omitted.

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